

7. molar heat of condensation

Part B True-False

8. AT 10. ST 12. NT
9. NT 11. NT

Part C Matching

13. a 15. b 17. d
14. e 16. c

Part D Questions and Problems

18. a. endothermic d. endothermic
b. endothermic e. exothermic
c. exothermic f. exothermic

$$19. \Delta H = 28.3 \frac{\text{g H}_2\text{O}(g)}{\text{mol H}_2\text{O}(g)} \times \frac{1 \text{ mol H}_2\text{O}(g)}{18.0 \text{ g H}_2\text{O}(g)} \\ \times \frac{6.01 \text{ kJ}}{1 \text{ mol H}_2\text{O}(g)} = 9.45 \text{ kJ}$$

$$20. \Delta H = 5.53 \frac{\text{mol NH}_4\text{NO}_3(g)}{\text{mol NH}_4\text{NO}_3(g)} \\ \times \frac{25.7 \text{ kJ}}{1 \text{ mol NH}_4\text{NO}_3(g)} = 142 \text{ kJ}$$

Section 17.4

Part A Completion

- sum
- enthalpy
- indirectly
- changed (reversed)
- standard heat of formation
- change
- one
- ΔH_f°
- zero
- subtracting

Part B True-False

11. NT 13. AT 15. NT
12. NT 14. NT

Part C Matching

16. b 18. a 20. c
17. c 19. d

Part D Questions and Problems



$$22. \Delta H_f^\circ (\text{products}) \\ = 3 \frac{\text{mol CO}_2(g)}{\text{mol CO}_2(g)} \times \frac{-393.5 \text{ kJ}}{1 \text{ mol CO}_2(g)} \\ = -1181 \text{ kJ} \\ \Delta H_f^\circ (\text{reactants}) \\ = 3 \frac{\text{mol CO}(g)}{\text{mol CO}(g)} \times \frac{-110.5 \text{ kJ}}{1 \text{ mol CO}(g)} + (-822.1 \text{ kJ}) \\ = -1154 \text{ kJ} \\ \Delta H^\circ = \Delta H_f^\circ (\text{products}) - \Delta H_f^\circ (\text{reactants}) \\ = -1181 \text{ kJ} - (-1154 \text{ kJ}) \\ = -27 \text{ kJ}$$

Practice Problems

Section 17.1

- $200.0 \text{ Cal} \times \frac{1000 \text{ cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ cal}} \times \frac{1 \text{ kJ}}{10^3 \text{ J}} \\ = 836.8 \text{ kJ}$
- $C = \frac{525.0 \text{ cal}}{25.0 \text{ g} \times 15.0^\circ\text{C}} = 1.40 \text{ cal/(g} \times ^\circ\text{C)}$
- $\Delta T = \frac{1255.0 \text{ J}}{100.0 \text{ g} \times 2.1 \text{ J/(g} \times ^\circ\text{C)}} = 6.0^\circ\text{C}$
- $q = 100.0 \text{ g} \times 120.0^\circ\text{C} \times 0.90 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \\ = 1.1 \times 10^4 \text{ J}$

Section 17.2

- $\Delta H = 150.0 \text{ g} \times 4.18 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \times 10^\circ\text{C} \\ = 6.3 \times 10^3 \text{ J} = 6.3 \text{ kJ}$
- $\Delta H = 15.0 \text{ g Ca(OH)}_2(g) \times \frac{1 \text{ mol Ca(OH)}_2(g)}{74.1 \text{ g Ca(OH)}_2(g)} \\ \times \frac{-65.2 \text{ kJ}}{1 \text{ mol Ca(OH)}_2(g)} = -13.2 \text{ kJ}$
- $\Delta H = 52.4 \text{ g CH}_4(g) \times \frac{1 \text{ mol CH}_4(g)}{16.0 \text{ g CH}_4(g)} \\ \times \frac{-890.2 \text{ kJ}}{1 \text{ mol CH}_4(g)} = 2.93 \times 10^3 \text{ kJ}$
- $4\text{NH}_3 + 5\text{O}_2(g) \rightarrow 4\text{NO}(g) + 6\text{H}_2\text{O}(g) \\ \Delta H = \frac{-226 \text{ kJ}}{1 \text{ mol NH}_3(g)} \times 4 \text{ mol NH}_3(g) \\ = -904 \text{ kJ}$

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Jeanette Adkins



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Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Prentice Hall Chemistry, 2000 *The Physics and Chemistry of Aqueous Ionic Solutions* M.C. Bellissent-Funel, G.W. Neilson, 2012-12-06 J E Enderby At the last NATO ASI on liquids held in Corsica August 1977 Professor de Gennes in his summary of that meeting suggested that the next ASI should concentrate on some specific aspect of the subject and mentioned explicitly ionic solutions as one possibility The challenge was taken up by Marie Claire Bellissent Funel and George Neilson I am sure that all the participants would wish to congratulate our two colleagues for putting together an outstanding programme of lectures round tables and poster session The theory which underlies the subject was covered by four leading authorities J P Hansen Paris set out the general framework in terms of the statistical mechanics of bulk and surface properties H L Friedman Stony Brook focused attention on ionic liquids at equilibrium and J B Hubbard considered non equilibrium properties such as the electrical conductivity and ionic friction coefficients Finally the basic theory of polyelectrolytes treated as charged linear polymers in aqueous solution was presented by J M Victor Paris *Computational Chemistry* Errol G. Lewars, 2024-06-04 This is the fourth edition of the successful textbook on computational chemistry which continues to provide a comprehensive introduction to the theory and practice of computational chemistry Notable updates include a review of references up to mid 2023 encompassing recent developments in scientific journals books and software The evolving prominence of density functional theory DFT is emphasized and attention is given to the increasing application of artificial intelligence in computational chemistry The book maintains key features from the previous edition delving into the mathematical intricacies of ab initio and density functional methods at an introductory level Clear explanations of matrix methods are provided offering a direct approach to obtaining energy levels and molecular orbitals Additionally each chapter includes sets of Easier and Harder drill questions with suggested answers at the end of the book enhancing the learning experience The book is intended for upper year undergraduate and graduate students studying computational and theoretical chemistry and for self study by researchers in universities and industry to whom computational chemistry may be useful *Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry* McMurry, Susan McMurry, 1990 **Study Guide and Solutions Manual for Organic Chemistry** Susan McMurry, John McMurry, 1988 Quantum Chemistry Walter Kauzmann, 2013-10-22 Quantum Chemistry An Introduction provides information pertinent to the fundamental aspects of quantum mechanics This book presents the theory of partial differentiation equations by using the classical theory of vibrations as a means of developing physical insight into this

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Physical and Chemical Equilibrium for Chemical Engineers Noel de Nevers, 2012-03-20 This book concentrates on the topic of physical and chemical equilibrium Using the simplest mathematics along with numerous numerical examples it accurately and rigorously covers physical and chemical equilibrium in depth and detail It continues to cover the topics found in the first edition however numerous updates have been made including Changes in naming and notation the first edition used the traditional names for the Gibbs Free Energy and for Partial Molal Properties this edition uses the more popular Gibbs Energy and Partial Molar Properties changes in symbols the first edition used the Lewis Randal fugacity rule and the popular symbol for the same quantity this edition only uses the popular notation and new problems have been added to the text Finally the second edition includes an appendix about the Bridgman table and its use

Bioinformatics Computing Bryan P. Bergeron, 2003 Comprehensive and concise this handbook has chapters on computing visualization large database designs advanced pattern matching and other key bioinformatics techniques It is a practical guide to computing in the growing field of Bioinformatics the study of how information is represented and transmitted in biological systems starting at the molecular level

Soil and Groundwater Remediation Chunlong Zhang, 2019-10-30 An introduction to the principles and

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