Equation Writing and Predicting Products

he treetesctions:

If you examine a car that has been in a junkyard for a while, you will notice that it has rusted. Rusting in a slow chemical reaction of the iron to the ear's body with oxygen gas. If a piece of sodium is put into water, a much more rapid chemical reaction occurs. Sodium reacts with water to produce sodium bydroxide and bydrogen gas. During this reaction, enough heat is liberated to ignite the hydrogen gas, causing it to explode.

Chemists observe what is happening in a chemical reaction and try to describe it in language that is simple and clear. A chemical equation uses formulas and symbols to describe the substances involved in a reaction, the physical state of the substance, the use of a catalyst, and relative proportions. The general form of an equation is:

Reactants ---- Preducts

In this investigation you will perform a series of reactions and make careful observations of the changes that occur. Using simple tests and your knowledge of chemistry, you will describe the identity of the products. With this information, you will write chemical equations to describe the reactions.

Pre-Eab Discussion:

Read firesigh the entire lab investigation and the relevant pages of your textbook. Then answer the questions that killow.

- 1. What constitutes a positive test for each of the following gases?
 2. expgen
 3. byptrogen
 4. water vapor
 6. structure (NFL)

 2. What is the proper way to smell a substance in the lab? Why should care be taken when smelling a gas such as ammonia?

 3. What is the role of a satisfyst in a reaction? How can yes tell when a substance serves as a catalyst? (This answer will be important later, spend some time on it now.)

 4. One way to identify limestone (CaCO₄) is to drop a small amount of hydrochionic acid (HCT) on it. A positive test results in a first of carbon districts being
 - produced. Write out in words the information represented by the following balanced element equation.

 CaCO₂(s) + 2 HCI (asp. —— CaCU₂(sq) + H₂O(l) + CO₂(g).

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