

CHEMISTRY SEMESTER 1 REVIEW QUESTIONS KEY

These questions are a small selection for each unit this semester. Refer to the exam scoring-guide and your self-reflection guide to determine which questions you should focus your studies on. Answers to select questions can be found at the end of this document.

Unit 1: Gas Laws

1. A gas is heated in a closed, rigid vessel. What happens to the molecules inside?
The molecules start to move faster. They hit the inside of the container with greater force, which increases the pressure.
2. A gas is compressed (volume decreased) at a constant temperature. What happens to the pressure inside the container? Explain your reasoning.
Since the volume is decreased, the molecules hit the sides of the container with increased frequency. This raises the pressure.
3. When inflating a car tire, the pressure doubles from 200 to 400 kPa. Why would the pressure double when the volume is increasing?
The volume may be increasing, but it is because more molecules are being added to the system. More molecules means more collisions, which increases the pressure.
4. What is the relation between temperature and pressure when volume is held constant?
With a constant volume, if temperature increases pressure increases as well. Think about heating an aerosol can (hair spray) – the volume is rigid and fixed. If you heat the can, the internal molecules move faster, hitting the sides harder and more frequently. This will increase the pressure so much that the can may explode.

Unit 2: Atomic structure and theory, periodic table

5. How many protons, neutron, and electrons are in a neutral atom of Ag-108?
Protons = 47
Neutrons = 61
Electrons = 47

6. Fill in the table below.

Name	Sodium	Nitrogen
Symbol	$^{23}_{11}\text{Na}$	$^{15}_7\text{N}$
Atomic number	11	7
Mass number	23	15
Protons	11	7
Neutrons	12	8
Electrons in a neutral atom	11	7

7. Analyze the following two isotopes. What is the same between both isotopes? What is different? $^{15}_7\text{N}$ and $^{14}_7\text{N}$
Both isotopes have the same number of protons (7) yet a different number of neutrons (8 and 7, respectively). The different number of neutrons yields a different mass number for the isotopes.
8. Chlorine has 2 isotopes, Cl-35 and Cl-37. Cl-35 has an abundance of 75.77% while Cl-37 has an abundance of 24.23%. Will the atomic mass of chlorine be closer to 35 or 37? Explain your reasoning?
Atomic mass will be closer to Cl-35 because 'atomic mass' refers to the average mass of all isotopes of an element. If most chlorine isotopes have a mass of 35 while only a few have a mass of 37, the average mass will be closer to 35.
9. Draw the Bohr model for sulfur. Label protons, neutrons, electrons, nucleus, and energy levels.



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Explorer Michael J. Padilla,2002 **Criminalistics: Forensic Science, Crime, and Terrorism** James E.

Girard,2013-11-13 Designed for students that are not biology chemistry or physics majors this fully revised and updated Third Edition of the best selling Criminalistics Forensic Science Crime and Terrorism provides a comprehensive introduction to forensic science the scientific principles that are the underpinnings of crime analysis and the practical application of these principles Essential topics such as fingerprint identification DNA ballistics detection of forgeries forensic toxicology computer forensics and the identification and analysis of illicit drugs are thoroughly explained in a reader friendly manner Unlike comparable texts the Third Edition includes coverage of important terrorism and homeland security issues including explosives cybercrime cyberterrorism and weapons of mass destruction The text is also the only book on the market with a detailed description of DNA and CODIS techniques used by professionals *Norman Hall's Asvab Preparation Book* Norman Hall,2015-01-02 Everything you need to know thoroughly covered in one book five ASVAB practice tests answer keys tips to boost scores military enlistment information study aids *Criminalistics* James E. Girard,James Girard,2011-01-28 Criminal

Investigations Forensic Science **Biophysical Chemistry: Techniques for the study of biological structure and function** Charles R. Cantor,Paul Reinhard Schimmel,1980 **Prentice Hall Exploring Life Science** Anthea Maton,1997

Physical Chemistry and Its Biological Applications Wallace Brey,2012-12-02 Physical Chemistry and Its Biological Applications presents the basic principles of physical chemistry and shows how the methods of physical chemistry are being applied to increase understanding of living systems Chapters 1 and 2 of the book discuss states of matter and solutions of nonelectrolytes Chapters 3 to 5 examine laws in thermodynamics and solutions of electrolytes Chapters 6 to 8 look at acid

base equilibria and the link between electromagnetic radiation and the structure of atoms Chapters 9 to 11 cover different types of bonding the rates of chemical reactions and the process of adsorption Chapters 12 to 14 present molecular aggregates magnetic resonance spectroscopy and photochemistry and radiation This book is useful to biological scientists for self study and reference With modest additions of mathematical material by the teacher the book should also be suitable for a full year major s course in physical chemistry

Science Books ,1972 *Biology* Leslie MacKenzie,David K. Arwine,Edward J. Shewan,Michael J. McHugh,2004-08 Originally developed by the Creation Research Society this classic text is now available in an updated and full color edition This hardbound text contains helpful questions and a thorough presentation of biology concepts Beautiful graphs and illustrations complement the text material that is scientifically accurate and true to six day young earth creationism Grades 9 10

Study Guide and Full Solutions Manual Susan McMurry,John McMurry,2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook

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Electrochemistry Christine Lefrou,Pierre Fabry,Jean-Claude Poignet,2012-05-24 This textbook offers original and new approaches to the teaching of electrochemical concepts principles and applications Throughout the text the authors provide a balanced coverage of the thermodynamic and kinetic processes at the heart of electrochemical systems The first half of the book outlines fundamental concepts appropriate to undergraduate students and the second half gives an in depth account of electrochemical systems suitable for experienced scientists and course lecturers Concepts are clearly explained and mathematical treatments are kept to a minimum or reported in appendices This book features Questions and answers for self assessment Basic and advanced level numerical descriptions Illustrated electrochemistry applications This book is accessible to both novice and experienced electrochemists and supports a deep understanding of the fundamental principles and laws of electrochemistry

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