

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the item.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure when the absolute temperature is 273 K.
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at 0°C and 1 atm.
5. A constant of nature that has values of 8.314 J/mol·K or 0.08206 L·atm/mol·K.
6. The a parameter of gas constant temperatures for determining the gas nonideality with pressure.
7. An experimental gas constant derived from the ideal gas law when absolute temperature is in kelvins.
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. Ratio of volume to mass

Column II

- a. standard pressure
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. Boyle's constant
- g. standard molar volume, pressure
- h. volume
- i. compressibility
- j. heat of fusion of water

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Subliming the solvent of a solid-solvent system gives a solid leaving a liquid solvent. Sublimation is used to separate a. the solid-solvent from the solid b. the solvent from the solid c. the solid from the solvent d. the solvent from the solid

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G Orfield



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Table of Contents Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers

1. Understanding the eBook Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers and Bestseller Lists
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 - Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers eBook Subscription Services

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- 6. Navigating Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers Compatibility with Devices
 - Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Interactive Elements Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
- 8. Staying Engaged with Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
- 9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Setting Reading Goals Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 8 Standardized Test Prep Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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