

Periodic Law and the Quantum Model

Use the PES spectrum of Phosphorus below to answer questions 1-3.



1. Which peak corresponds to the 1s orbital?
(A) 1.06
(B) 1.95
(C) 13.5
(D) 18.7
(E) 208
2. Which peak corresponds to the valence 3p orbital?
(A) 1.06
(B) 1.95
(C) 13.5
(D) 18.7
(E) 208
3. What would you predict for the intensity and binding energy for the 3p orbital for that of sulfur?
(A) Higher intensity, higher binding energy
(B) Higher intensity, lower binding energy
(C) Lower intensity, higher binding energy
(D) Lower intensity, lower binding energy

Psi Ap Chemistry Periodic Trends Test

Judd E. Hollander



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