

12**STOICHIOMETRY****SECTION 12.1 THE ARITHMETIC OF EQUATIONS** (Pages 353–356)

This section emphasizes how to calculate the amounts of reactants required or products formed in a stoichiometric process. As you learn how to calculate mole-to-mole ratios from balanced equations, let's recall what we learned in Chapter 10: moles, molar mass, and Avogadro's number.

► Using Everyday Equations (Pages 353–355)

- How can you determine the quantities of reactants and products in a chemical reaction?

You can either: **A. Balance the equation**

- Quantity usually means the **amount** of a substance expressed in grams or moles.

- A balanced equation shows the relative amounts of reactants (PbO, 2 mol PbO), 1 mole O₂, and 2 moles Pb. This is "balancing" because it conserves mass and mole relationships.

**► Using Balanced Chemical Equations** (Page 354)

- To determine the amount of product formed by a balanced equation, multiply the balanced equation by the **mole ratio**.
- Calculations using balanced equations are called **mole calculations**.

► Interpreting Chemical Equations (Pages 356–357)

- What effect does each term in a chemical equation have on a result?

For example, a coefficient can be multiplied out and signs should be changed if the balanced equation is rearranged.

- Consider the last two columns above. What do these terms tell us about stoichiometry? Does it affect your results?

a. **reactants**

b. **products**

c. **moles**

d. **numbers of particles**

e. **sides-of particles**

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