

## 9-3 Review and Reinforcement

### Classifying chemical Reactions

On the line on the left, write the letter of the type of chemical reaction represented by each equation below.

- |  |                                |
|--|--------------------------------|
| _____ 1. $A + B \rightarrow AB$        | a. decomposition reaction      |
| _____ 2. $A + BX \rightarrow AX + B$   | b. direct combination reaction |
| _____ 3. $AX + BY \rightarrow AY + BX$ | c. double-replacement reaction |
| _____ 4. $AB \rightarrow A + B$        | d. single-replacement reaction |

Describe in words each of the following types of reactions.

5. direct combination reaction

\_\_\_\_\_

6. decomposition reaction

\_\_\_\_\_

7. single-replacement reaction

\_\_\_\_\_

8. double-replacement reaction

\_\_\_\_\_

Decide whether each of the following equations represents a direct combination (dc), decomposition (d), single-replacement (sr), or double-replacement (dr) reaction. Write your answer on the line.

- \_\_\_\_\_ 9.  $\text{CO}_2 \rightarrow \text{C} + \text{O}_2$
- \_\_\_\_\_ 10.  $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}$
- \_\_\_\_\_ 11.  $\text{S}_8 + 4\text{Cl} \rightarrow 8\text{SCl}_2$
- \_\_\_\_\_ 12.  $\text{BaCl}_2 + 2\text{NaOH} \rightarrow 2\text{NaCl} + \text{Ba(OH)}_2$
- \_\_\_\_\_ 13.  $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
- \_\_\_\_\_ 14.  $\text{CH}_4 \rightarrow \text{C} + 2\text{H}_2$
- \_\_\_\_\_ 15.  $\text{Pb(NO}_3)_2 + \text{Mg} \rightarrow \text{Pb} + \text{Mg(NO}_3)_2$

## Section 2 Reinforcement Rates Of Chemical Reactions

**Mr. Rohit Manglik**



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