

Name: ANSWERS150

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question. (20 marks)

1. What is the purpose of a salt bridge in a galvanic cell?
a. to balance charge
b. ☒ to complete the circuit
c. to absorb electrons
d. to release electrons
2. Which is the correct net ionic equation for the following reaction $\text{MgSO}_4(\text{aq}) + \text{Cu(s)} \rightarrow \text{Mg(s)} + \text{CuSO}_4(\text{aq})$?
a. ☒ $\text{Mg}^{2+}(\text{aq}) + \text{Cu(s)} \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Mg(s)}$
b. $\text{MgSO}_4(\text{aq}) + \text{Cu(s)} \rightarrow \text{Mg(s)} + \text{CuSO}_4(\text{aq})$
c. $\text{Mg}^{2+}(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq})$
d. $\text{Mg}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{Cu(s)} \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{Mg(s)}$
3. Identify the oxidizing agent in the reaction: $\text{Sn}^{2+}(\text{aq}) + \text{Cl}_2(\text{g}) \rightarrow \text{Sn}^{4+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq})$
a. $\text{Sn}^{2+}(\text{aq})$
b. ☒ $\text{Cl}_2(\text{g})$
c. $\text{Sn}^{4+}(\text{aq})$
d. $\text{Sn}^{4+}(\text{aq})$
4. Identify the reducing agent in the reaction: $2\text{Fe(s)} + \text{O}_2(\text{g}) + 2\text{H}_2\text{O(l)} \rightarrow 2\text{Fe(OH)}_2(\text{s})$
a. $\text{H}_2\text{O(l)}$
b. ☒ Fe(s)
c. Fe(s) and $\text{O}_2(\text{g})$
d. $\text{O}_2(\text{g})$
5. In the reaction: $\text{Al(s)} + \text{O}_2(\text{g}) \rightarrow \text{Al}_2\text{O}_3(\text{s})$, how many electrons does an atom of aluminum lose or gain?
a. aluminum loses 1 electron
b. aluminum gains 1 electron
c. ☒ aluminum loses 3 electrons
d. aluminum gains 3 electrons
6. Which of the following is a stronger reducing agent than Mg(s) ?
a. ☒ Ca(s)
b. Pb(s)
c. Sn(s)
d. Fe(s)
7. Which of these reactions will proceed spontaneously, based on relative strength of reducing agents?
a. ☒ $\text{Zn(s)} + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{Zn}^{2+}(\text{aq}) + \text{Cu(s)}$
b. $\text{Sn(s)} + \text{Cr}^{3+}(\text{aq}) \rightarrow \text{Sn}^{2+}(\text{aq}) + \text{Cr(s)}$
c. $\text{Cd(s)} + \text{Al}^{3+}(\text{aq}) \rightarrow \text{Cd}^{2+}(\text{aq}) + \text{Al(s)}$
d. $\text{Ag(s)} + \text{Zn}^{2+}(\text{aq}) \rightarrow \text{Ag}^{+}(\text{aq}) + \text{Zn(s)}$
8. What must be added to the following half-reaction for it to be balanced? $\text{Al}^{3+}(\text{aq}) \rightarrow \text{Al(s)}$
a. 1 electron to the left side
b. ☒ 3 electrons to the left side
c. 1 electron to the right side
d. 3 electrons to the right side
9. What must be added to the following half-reaction for it to be balanced? $10\text{H}^{+} + \text{NO}_2^{-}(\text{aq}) + 8\text{e}^{-} \rightarrow \text{NH}_4^{+}(\text{aq}) + 3\text{H}_2\text{O}$
a. 3 water molecules to the left side
b. ☒ 3 water molecules on the right side
c. 1 water molecule to the left side
d. 1 water molecule to the right side
10. How many hydrogen ions must be added to the half-reaction to balance the mass of hydrogen atoms?
 $\text{H}_2\text{O} + \text{CN}^{-}(\text{aq}) \rightarrow \text{CNO}^{-}(\text{s}) + \underline{2}\text{H}^{+}$
a. 1 on the left side
b. ☒ 2 on the left side
c. 2 on the right side
d. 1 on the right side
11. Which of the following reactions is an example of a disproportionation reaction?
a. $\text{PbCl}_2(\text{s}) \rightarrow \text{Pb(s)} + \text{Cl}_2(\text{g})$
b. $\text{Fe(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{FeSO}_4(\text{aq}) + \text{Cu(s)}$
c. $\text{NO}_2^{-}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightarrow \text{HNO}_2(\text{g})$
d. ☒ $2\text{Cl}_2(\text{g}) + 6\text{OH}^{-}(\text{aq}) \rightarrow 5\text{Cl}^{-}(\text{aq}) + \text{ClO}_2^{-}(\text{aq}) + 3\text{H}_2\text{O(l)}$
one reactant + element is both oxidized + reduced.

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