

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                    |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution   |
| 4. autoionization | 5. acid rain                | 6. buffer solution |
|                   | 7. acid solution            | 8. pH              |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the spaces provided. Be sure to include units when appropriate.

1. What substance is \_\_\_\_\_, which hydrogen ion concentration is \_\_\_\_\_? The concentration of these acids present in rain is \_\_\_\_\_.
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## Section 19 2 Hydrogen And Ions Acidity Answer Key

**Barbara Janson Cohen, Kerry L. Hull**



## Section 19 2 Hydrogen And Ions Acidity Answer Key:

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exam file resource containing a complete set of exam style questions in a format that can be used throughout Years 10 and 11 or as a resource for a revision programme is included      *Ophthalmology, E-Book* Myron Yanoff, Jay S. Duker, 2022-11-06

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