

# Relative Strengths of **Acids** and **Bases** in Aqueous Solutions at 25°C

Name of Acid	Acid		Base	$K_a$
Hydriodic	HI	→	$H^+ + I^-$	$>> 1$
Hydrobromic	HBr	→	$H^+ + Br^-$	$>> 1$
Hydrochloric	HCl	→	$H^+ + Cl^-$	$>> 1$
Sulfuric	$H_2SO_4$	→	$H^+ + HSO_4^-$	$> 1$
Nitric	$HNO_3$	→	$H^+ + NO_3^-$	$> 1$
Hydronium Ion	$H_3O^+$	⇌	$H^+ + H_2O$	$5.50 \times 10^{-1}$
Oxalic	$H_2C_2O_4$	⇌	$H^+ + HC_2O_4^-$	$5.90 \times 10^{-2}$
Sulfurous	$H_2SO_3$	⇌	$H^+ + HSO_3^-$	$1.41 \times 10^{-2}$
Hydrogen Sulfate Ion	$HSO_4^-$	⇌	$H^+ + SO_4^{2-}$	$1.20 \times 10^{-2}$
Phosphoric	$H_3PO_4$	⇌	$H^+ + H_2PO_4^-$	$7.52 \times 10^{-3}$
Nitrous	$HNO_2$	⇌	$H^+ + NO_2^-$	$5.62 \times 10^{-4}$
Hydrofluoric	HF	⇌	$H^+ + F^-$	$3.53 \times 10^{-4}$
Formic	$HCO_2H$	⇌	$H^+ + HCO_2^-$	$1.78 \times 10^{-4}$
Benzoic	$C_6H_5CO_2H$	⇌	$H^+ + C_6H_5CO_2^-$	$6.46 \times 10^{-5}$
Acetic	$CH_3CO_2H$	⇌	$H^+ + CH_3CO_2^-$	$1.76 \times 10^{-5}$
Carbonic	$H_2CO_3$	⇌	$H^+ + HCO_3^-$	$4.30 \times 10^{-7}$
Hydrogen Sulfite Ion	$HSO_3^-$	⇌	$H^+ + SO_3^{2-}$	$1.02 \times 10^{-7}$
Dihydrogen Phosphate Ion	$H_2PO_4^-$	⇌	$H^+ + HPO_4^{2-}$	$6.23 \times 10^{-8}$
Boric	$H_3BO_3$	⇌	$H^+ + H_2BO_3^-$	$5.79 \times 10^{-10}$
Ammonium Ion	$NH_4^+$	⇌	$H^+ + NH_3$	$5.64 \times 10^{-10}$
Hydrogen Carbonate Ion	$HCO_3^-$	⇌	$H^+ + CO_3^{2-}$	$5.61 \times 10^{-11}$
Monohydrogen Phosphate Ion	$HPO_4^{2-}$	⇌	$H^+ + PO_4^{3-}$	$2.20 \times 10^{-13}$
Water	$H_2O$	⇌	$H^+ + OH^-$	$1.00 \times 10^{-14}$

Strong Acids

Weak Acids

Weakest Acids

Weakest Bases

Weak Bases

Stronger Bases

All reactions assume water as a reactant.



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$$K_a = [H^+][A^-] / [HA]$$

$$pK_a = -\log K_a$$

$$pH = -\log [H^+]$$

## Section 2 Study Guid Strength Of Acid And Bases

**Leo J. Malone**



## Section 2 Study Guid Strength Of Acid And Bases:

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03  
The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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*Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition* John McMurry, Susan McMurry, 1999. Provides worked-out solutions to text problems along with chapter-by-chapter outlines and a variety of self-tests at the end of each chapter.

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**Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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**Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson** Norman E. Griswold, Henry Fuller Holtzclaw,1988

Organic Chemistry, Student Study Guide and Solutions Manual David R. Klein,2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Existing textbooks provide extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems

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*Basic Concepts of Chemistry, Study Guide* Leo J. Malone, 2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of

the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems

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