

## 1. Identify the attractive forces associated with...

## a. Ionic Bonds:

Attractive electrostatic forces between negative charged ion and adjacent positively charged ion

## b. Covalent Bonds:

Electrostatic forces between negatively charged electrons and adjacent positively charged nucleus

## 2. Identify two similarities between ionic and covalent bonds

- Both involve valence electron clouds
- Both involve electrostatic attractions between oppositely charged species
- The formation of bond begins with the valence electrons of two atoms experiencing the attractive force of adjacent positive nuclei
- Both Ionic and Covalent bonds are strong

## 3. Identify two difference between ionic and covalent bonds

- Ionic Bonds involve the transfer of valence electrons forming ions whereas covalent bonds involve the sharing of valence electrons with no ion formation
- Ionic bonds form only between metals and non-metals whereas covalent bonds usually (but not always) form between two non-metals
- Ionic bonding does not result in molecule formation whereas covalent bonding usually does
- The attractive forces associated with covalent bonds are the electrostatic forces between negatively charged electrons and adjacent positively charged nuclei

## 4. Complete the following table by writing formulas of the compounds formed from the pairs of elements.

Elements	Compound Formula	$\Delta EN$ Value	Nature of Bonds Present
(a) rubidium and oxygen	$Rb_2O$	2.7	ionic
(b) strontium and bromine	$SrBr_2$	1.8	ionic
(c) carbon and sulphur	$CS_2$	0	covalent
(d) silicon and chlorine	$SiCl_4$	1.2	polar covalent

5. Magnesium is a metal and sulphur is a non-metal. Compare the  $\Delta EN$  value for these elements in  $MgS$  to  $\Delta EN$  value for the non-metal hydrogen and oxygen in water ( $H_2O$ ). Which of the two compounds possesses a greater amount of ionic character in its bond?

$\Delta EN$  for  $MgS$  = 1.3

$\Delta EN$  for  $H_2O$  = 1.4

We see that the bonds in water actually possess slightly more ionic character than those in magnesium sulphide even though the former compound contains two non-metals and the latter compound contains a metal and a non-metal

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**CL Gary**



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