

Section 12.3 Limiting Reagent & Percent Yield

Answer Key

Michelle Ganian

Vocabulary

1. e
2. c
3. d
4. a
5. b

1. Answer: 127g Cu
Because...

Find the limiting reagent:

$$54\text{g Al} \times 1 \text{ mole Al}/27\text{g Al} = 2 \text{ mole Al}$$

$$319\text{g CuSO}_4 \times 1 \text{ mole CuSO}_4/159.5\text{g CuSO}_4 = 2 \text{ mole CuSO}_4 \text{ (limiting reagent)}$$

$$2 \text{ mole CuSO}_4 \times 3 \text{ mole Cu}/3 \text{ mole CuSO}_4 \times 63.5\text{g Cu}/1 \text{ mole Cu} = 127\text{g Cu}$$

(Multiply the founded limiting reagent with moles of Copper (3) over moles of CuSO_4 (3) from the equation to cancel out moles of CuSO_4 .)

Then, to cancel out 3 moles Cu, multiply it with grams Cu (because that is also what the question asks for) over 1 mole Cu. After canceling and multiplying out what is left, you come to 127g Cu.)

2. $27\text{g Cu} \times 1 \text{ mole Cu}/63.5\text{g Cu} \times 2 \text{ mole Cr}/3 \text{ mole Cu} \times 52\text{g Cr}/1 \text{ mole Cr} = 14.74\text{g Cr}$

("excess of CuSO_4 " is irrelevant to the problem you are solving, so you can cross that out. You are left only to work with 27g Cu, and need to find grams Cr. So you divide 27g Cu by its molar mass [from periodic table] and then multiply it by the mole to mole ratio of 2 moles Cr over 3 moles Cu. You put 3 moles Cu on the bottom to cancel it out with 1 mole Cu. Lastly, to cancel out 2 moles Cr, multiply it with the molar mass [52g] over 1 mole Cr and multiply out what is uncrossed.)

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**Dana W. Mayo, Ronald M. Pike, David C.
Forbes**

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