

Molality, *continued* ▼

- The molality of any solution can be calculated by dividing the number of moles of solute by the number of kilograms of solvent: ▼

$$\text{molality } (m) = \frac{\text{amount of solute (mol)}}{\text{mass of solvent (kg)}} \quad \blacktriangledown$$

- Unlike molarity, which is a ratio of which the denominator is liters of *solution*, molality is per *kilograms of solvent*. ▼
 - Molality is used when studying properties of solutions related to vapor pressure and temperature changes, because molality does not change with temperature.



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Holt Chemistry Ralph Thomas Myers, 2004 Ebook: Chemistry: The Molecular Nature of Matter and Change
Silberberg, 2015-01-16 Ebook *Chemistry The Molecular Nature of Matter and Change* **Holt Chemistry** Holt Rinehart & Winston, 2003-01-24 *Chemistry3* Andrew Burrows, John Holman, Andrew Parsons, Gwen Pilling, Gareth Price, 2017
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