

## Part 2

# GCSE Chemistry OCR A (Combined Science - **HIGHER**) Paper 2 Exam Walkthrough

### Combined Science Higher - Chemistry

### Walkthrough Mark Paper

12  
(5)

Look at the equation for the equilibrium reaction between sulfur dioxide,  $\text{SO}_2$ , oxygen and sulfur trioxide,  $\text{SO}_3$ , in a closed system.



- (a) Predict the effect of adding more oxygen to the equilibrium mixture.

Explain your answer.

(2)

- (b) Predict the effect of increasing the total pressure in the equilibrium mixture.

Explain your answer.

(2)

- (c) Predict the effect of increasing the **temperature** of the equilibrium mixture.

Explain your answer.

(2)

- (d) Sulfur dioxide for this reaction is made by burning sulfur.



Calculate the mass of sulfur needed to make 48 tonnes of sulfur dioxide.

Give your answer to 2 significant figures.

The relative atomic mass,  $A_r$ , of S is 32.1 and of O is 16.0.

# Ocr Past Paper 2013 June Chemistry

**Rosina Ehmann**



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### **Table of Contents Ocr Past Paper 2013 June Chemistry**

1. Understanding the eBook Ocr Past Paper 2013 June Chemistry
  - The Rise of Digital Reading Ocr Past Paper 2013 June Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Past Paper 2013 June Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Past Paper 2013 June Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Past Paper 2013 June Chemistry
  - Personalized Recommendations
  - Ocr Past Paper 2013 June Chemistry User Reviews and Ratings

- Ocr Past Paper 2013 June Chemistry and Bestseller Lists
- 5. Accessing Ocr Past Paper 2013 June Chemistry Free and Paid eBooks
  - Ocr Past Paper 2013 June Chemistry Public Domain eBooks
  - Ocr Past Paper 2013 June Chemistry eBook Subscription Services
  - Ocr Past Paper 2013 June Chemistry Budget-Friendly Options
- 6. Navigating Ocr Past Paper 2013 June Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr Past Paper 2013 June Chemistry Compatibility with Devices
  - Ocr Past Paper 2013 June Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Past Paper 2013 June Chemistry
  - Highlighting and Note-Taking Ocr Past Paper 2013 June Chemistry
  - Interactive Elements Ocr Past Paper 2013 June Chemistry
- 8. Staying Engaged with Ocr Past Paper 2013 June Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Past Paper 2013 June Chemistry
- 9. Balancing eBooks and Physical Books Ocr Past Paper 2013 June Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Past Paper 2013 June Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Past Paper 2013 June Chemistry
  - Setting Reading Goals Ocr Past Paper 2013 June Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Past Paper 2013 June Chemistry
  - Fact-Checking eBook Content of Ocr Past Paper 2013 June Chemistry
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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