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The Hydrogen Atom - Student Guide

Background Material

Carefully read the background pages entitled Energy Levels, Light, and Transitions and answer the following questions to check your understanding.

Question 1: Complete the following table which compares how the Bohr Model and the Quantum model represent the Hydrogen atom. In some cases they both describe things in the same way and in some cases they do not.

Concept	Bohr Model	Quantum Model	
Nucleus	positively charged proton at the center	positively charged proton at the center	
Electrons	Negatively charged cloud	negatively charged cloud	
Orbital	Electrons travel in a circular orbits around nucleus	Electrons are in a certain position in orbit	
Location of electron in the first orbital	always located at the Bohr radius	Always located on electron cloud	
Energy of the Energy Levels	80-8642	$E = -E_0/rr^2$	

Question 2: Complete the following table which relates the parameters of two different photons by circling the appropriate relationship. The first row is completed for you: "A red photon has a larger wavelength, smaller frequency, smaller energy, and the same velocity through space as a blue photon".

Satpbiology Student Review Guide Answer Key

Daniela Niemeyer

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