

16. Increasing the temperature generally.....
- decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
 - increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
 - increases the solubility of both a solid and a gas in a liquid.
 - decreases the solubility of both a solid and a gas in a liquid.
17. A white precipitate was formed when $\text{BaCl}_{2(aq)}$ was added to an aqueous solution of a salt X. The precipitate dissolved in dilute HCl with rapid effervescence. Salt X is likely to contain
- SO_4^{2-} ions.
 - SO_3^{2-} ions.
 - NO_3^- ions.
 - CO_3^{2-} ions.
18. Before a reaction could take place, there should be
- ionization of reactant particles.
 - breakage of bonds of products.
 - breakage of bonds of reactants.
 - ionization of product particles.
19. Consider the following reaction equation:
- $$\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$$
- Silicon(IV) oxide is acting as
- a basic oxide.
 - a reducing agent.
 - an acidic oxide.
 - an oxidizing agent.
20. Which of the following acids would form normal salt only?
- Tetraoxosulphate(VI) acid
 - Trioxonitrate(V) acid
 - Tetraoxophosphate(V) acid
 - Trioxosulphate(IV) acid
21. What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing 100 cm^3 of oxygen and 900 cm^3 of nitrogen gas?
- 1.0 atm
 - 0.1 atm
 - 9.0 atm
 - 0.9 atm
22. Graphite is used as a dry lubricant due to the presence of
- mobile electrons.
 - free electrons.
 - octahedral structures.
 - layered structures.
23. Which of the following gases is alkaline?
- NO_2
 - CO_2
 - NH_3
 - N_2O
24. Which of the following statements about an equilibrium system is correct?
- Forward and backward reactions occur at the same rate.
 - The concentrations of reactants must equal that of the products.
 - The concentrations of reactants and products can be changed by adding a catalyst.
 - The concentrations of reactants and products are not affected by a change in temperature.

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Table of Contents Today Waec Question On Chemistry Paper

1. Understanding the eBook Today Waec Question On Chemistry Paper
 - The Rise of Digital Reading Today Waec Question On Chemistry Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Today Waec Question On Chemistry Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Today Waec Question On Chemistry Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Today Waec Question On Chemistry Paper
 - Personalized Recommendations
 - Today Waec Question On Chemistry Paper User Reviews and Ratings
 - Today Waec Question On Chemistry Paper and Bestseller Lists
5. Accessing Today Waec Question On Chemistry Paper Free and Paid eBooks

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 - Today Waec Question On Chemistry Paper eBook Subscription Services
 - Today Waec Question On Chemistry Paper Budget-Friendly Options
6. Navigating Today Waec Question On Chemistry Paper eBook Formats
- ePub, PDF, MOBI, and More
 - Today Waec Question On Chemistry Paper Compatibility with Devices
 - Today Waec Question On Chemistry Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Today Waec Question On Chemistry Paper
 - Highlighting and Note-Taking Today Waec Question On Chemistry Paper
 - Interactive Elements Today Waec Question On Chemistry Paper
8. Staying Engaged with Today Waec Question On Chemistry Paper
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Today Waec Question On Chemistry Paper
9. Balancing eBooks and Physical Books Today Waec Question On Chemistry Paper
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Today Waec Question On Chemistry Paper
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Today Waec Question On Chemistry Paper
- Setting Reading Goals Today Waec Question On Chemistry Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Today Waec Question On Chemistry Paper
- Fact-Checking eBook Content of Today Waec Question On Chemistry Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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