

# Chemical Basis of Life

## Vocabulary

Polarity

Concentration

4 carbon compounds

monosaccharide

lipid

enzyme

Hydrogen bond

Solution

monomer

polysaccharide

fatty acid

catalyst

Adhesion

Solute

polymer

amino acid

nucleic acid

pH

Cohesion

Solvent

carbohydrate

protein

nucleotide

specific heat

## Concepts to Know

### WATER

#### Why so special and important?

Nearly 75% of our earth is covered in water.

Our bodies are 60-65% water. Our blood is 85% water.

None of the many reactions that occur in our bodies would occur without water.

Composition of water:  $\text{H}_2\text{O}$

#### Properties of Water

- Water is **POLAR**.

The diagram shows that one side of the water molecule is **positively charged** and the other is **negatively charged**.

The molecule itself is neutral. These opposite charges make water a **polar molecule**.

- Water undergoes **ADHESION** and **COHESION** due to its polarity.

The **negative oxygen** of one water molecule is attracted to the **positive hydrogen** of another molecule forming a **hydrogen bond**. In other words water likes to stick to itself. Water sticking to water is called **cohesion**. Water's polarity allows it to attract to other surfaces too (like glass). Water sticking to something else is called **adhesion**.

- In the diagram to the right use dotted lines to draw in the bonds that form between water molecules.

**Positives connect to negatives, up to four bonds per water molecule.**

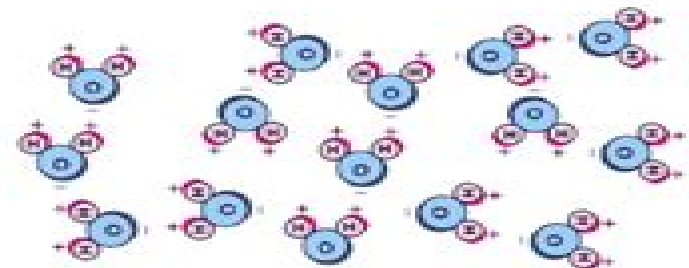
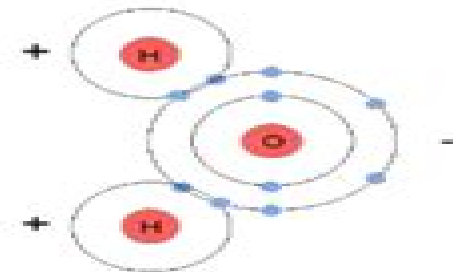
- What is the name of this type of bond?

**Hydrogen or "H" bond**

- What special property do the bonds give to water.

**Several: cohesion first and foremost (water sticking to itself).**

**Creates surface tension on the water (water skipper can seemingly "walk" on water)**



# The Chemical Basis Of Life Chapter 3 Packet Answers

**Yi-Tong Ma**



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## **Table of Contents The Chemical Basis Of Life Chapter 3 Packet Answers**

1. Understanding the eBook The Chemical Basis Of Life Chapter 3 Packet Answers
  - The Rise of Digital Reading The Chemical Basis Of Life Chapter 3 Packet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying The Chemical Basis Of Life Chapter 3 Packet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an The Chemical Basis Of Life Chapter 3 Packet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from The Chemical Basis Of Life Chapter 3 Packet Answers
  - Personalized Recommendations
  - The Chemical Basis Of Life Chapter 3 Packet Answers User Reviews and Ratings
  - The Chemical Basis Of Life Chapter 3 Packet Answers and Bestseller Lists
5. Accessing The Chemical Basis Of Life Chapter 3 Packet Answers Free and Paid eBooks
  - The Chemical Basis Of Life Chapter 3 Packet Answers Public Domain eBooks
  - The Chemical Basis Of Life Chapter 3 Packet Answers eBook Subscription Services
  - The Chemical Basis Of Life Chapter 3 Packet Answers Budget-Friendly Options
6. Navigating The Chemical Basis Of Life Chapter 3 Packet Answers eBook Formats

- ePub, PDF, MOBI, and More
  - The Chemical Basis Of Life Chapter 3 Packet Answers Compatibility with Devices
  - The Chemical Basis Of Life Chapter 3 Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of The Chemical Basis Of Life Chapter 3 Packet Answers
    - Highlighting and Note-Taking The Chemical Basis Of Life Chapter 3 Packet Answers
    - Interactive Elements The Chemical Basis Of Life Chapter 3 Packet Answers
  8. Staying Engaged with The Chemical Basis Of Life Chapter 3 Packet Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers The Chemical Basis Of Life Chapter 3 Packet Answers
  9. Balancing eBooks and Physical Books The Chemical Basis Of Life Chapter 3 Packet Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection The Chemical Basis Of Life Chapter 3 Packet Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine The Chemical Basis Of Life Chapter 3 Packet Answers
    - Setting Reading Goals The Chemical Basis Of Life Chapter 3 Packet Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of The Chemical Basis Of Life Chapter 3 Packet Answers
    - Fact-Checking eBook Content of The Chemical Basis Of Life Chapter 3 Packet Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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