

Periodic Law

OBJECTIVES

- To observe and investigate the periodic trends for reactivity, density, pH and solubility within groups 1, 2 and 14 in the periodic table of the elements.

INTRODUCTION

When the elements are arranged in order of increasing atomic number, there is a periodic occurrence of properties that leads to the grouping of elements on the periodic table. Elements in vertical columns form groups or families that are characterized by marked similarities in physical and chemical properties. These similarities are due to the fact that all of the elements within a group have the same number of valence electrons. Within a group, periodic trends in certain properties can be observed. The properties of density and solubility, for example, are known to vary regularly within groups of elements.

PROCEDURE

PART A

The teacher will demonstrate the reactivity of three elements (Lithium, Sodium and Potassium) in group 1, by adding a small piece of the metal to water. Observe data sheet and record observations and rank the three elements in order of reactivity.

Element	Reactivity	Observations
Li		
Na		
K		

PART B

- Pour about 5 mL of distilled water into a clean dry test tube.
- FRONT LAB TABLE:** Remove the pH probe with caution from the water bottle and insert the pH probe into the distilled water. Note the pH on the "Lab Quest". Remove the pH probe and store with the water bottle.
- FRONT LAB TABLE:** Add a few drops of phenolphthalein solution to the reaction tube. Phenolphthalein solution turns pink in the presence of a base.

The Periodic Law Lab Answers

DP Hallahan



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