Angular Quantum Number (5)

- also known as asmulfied quantum number
- describes the shapes of the orbitols
- Orbifolis howe shapes that are best described as spherical (F = 0), polar (F = 1), or cloverlead (F = 2). They can even take an more complex shapes as the value of the angular quantum number becomes larger.



Magnetic Guantum Number (m)

- to describe the orientation in space of a particular orbital.
- It is called the magnetic quantum number because the effect of different orientations of orbitals was first observed in the presence of a magnetic field.

Bules Governing the Albrared Combinations of Quantum Numbers

- The free quartum numbers (n. i. and m) that describe an arbital are integers 0, 1, 2, 3, and so on.
- 2. The principal quantum number [1] cannot be sen. The allowed values of n are therefore 1.
- 2. 3. 4. and so on.
- 3. The angular quantum number $|0\rangle$ can be any integer between 0 and n-1. If n=3, for example, l can be either 0, 1, or 2.
- The magnetic quantum number (m) can be any integer between if and +i. If i + 2, m can be either -2, -1, 0, +1, or +2.

Attornec Orbitals

A complete control of the control of

 The second character identifies the subshell. By convention, the following lowercase letters are used to indicate different subshells.

SC	1-0
pc.	$\ell=3$
alt:	$\ell = 2$
p.	$\ell = 3$

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James E. Brady, Fred Senese

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