

# REPORT SHEET

## Rates of Chemical Reaction

Name \_\_\_\_\_

Date \_\_\_\_\_

Section \_\_\_\_\_

Partner's Name \_\_\_\_\_

Instructor \_\_\_\_\_

(1.5 pts. total for Tables 1-3)

Table 1, Experiment #1

Initial moles of $S_2O_8^{2-}$		Cumulative time recorded from stopwatch		Cumulative time converted to seconds (s)	Mole of $S_2O_8^{2-}$ Consumed	Mole of $S_2O_8^{2-}$ Remaining	Concentration of $S_2O_8^{2-}$ Remaining (mol/L)
Aliquot #		(min)	(sec)				
1		1 min	46 sec	106	$1.0 \times 10^{-4} \text{ mol}$	$0.0025 \text{ mol}$	$0.046 \text{ M}$
2		3	24 sec	204	$2.0 \times 10^{-4} \text{ mol}$	$0.0023 \text{ mol}$	$0.043 \text{ M}$
3		4	29 sec	269	$3.0 \times 10^{-4} \text{ mol}$	$0.0021 \text{ mol}$	$0.040 \text{ M}$
4		5	48 sec	345	$4.0 \times 10^{-4} \text{ mol}$	$0.0020 \text{ mol}$	$0.038 \text{ M}$
5		7	03	423	$5.0 \times 10^{-4} \text{ mol}$	$0.0019 \text{ mol}$	$0.035 \text{ M}$

Table 2, Experiment #2

Initial moles of $S_2O_8^{2-}$		Cumulative time recorded from stopwatch		Cumulative time converted to seconds (s)	Mole of $S_2O_8^{2-}$ Consumed	Mole of $S_2O_8^{2-}$ Remaining	Concentration of $S_2O_8^{2-}$ Remaining (mol/L)
Aliquot #		(min)	(sec)				
1		2	46	165	$1.0 \times 10^{-4} \text{ mol}$	$0.0023 \text{ mol}$	$0.046 \text{ M}$
2		5	19	319	$2.0 \times 10^{-4} \text{ mol}$	$0.0022 \text{ mol}$	$0.043 \text{ M}$
3		8	33	513	$3.0 \times 10^{-4} \text{ mol}$	$0.0021 \text{ mol}$	$0.040 \text{ M}$
4		11	11	671	$4.0 \times 10^{-4} \text{ mol}$	$0.0020 \text{ mol}$	$0.038 \text{ M}$
5		13	17	797	$5.0 \times 10^{-4} \text{ mol}$	$0.0019 \text{ mol}$	$0.035 \text{ M}$

Table 3, Experiment #3

Initial moles of $S_2O_8^{2-}$		Cumulative time recorded from stopwatch		Cumulative time converted to seconds (s)	Mole of $S_2O_8^{2-}$ Consumed	Mole of $S_2O_8^{2-}$ Remaining	Concentration of $S_2O_8^{2-}$ Remaining (mol/L)
Aliquot #		(min)	(sec)				
1		1	20	80	$1.0 \times 10^{-4} \text{ mol}$	$0.0042 \text{ mol}$	$0.0074 \text{ M}$
2		2	30	150	$2.0 \times 10^{-4} \text{ mol}$	$0.0046 \text{ mol}$	$0.0085 \text{ M}$
3		3	43	223	$3.0 \times 10^{-4} \text{ mol}$	$0.0045 \text{ mol}$	$0.0087 \text{ M}$
4		3	59	233	$4.0 \times 10^{-4} \text{ mol}$	$0.0044 \text{ mol}$	$0.0083 \text{ M}$
5		5	06	306	$5.0 \times 10^{-4} \text{ mol}$	$0.0043 \text{ mol}$	$0.0080 \text{ M}$

# Rates Of Chemical Reactions Lab Answers

**Hongru Du**



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