

Name _____
 Period ____ Date ____/____/____

Chemical Kinetics

RATE LAWS

1. Consider the reaction: $2 \text{NO(g)} + \text{O}_2\text{(g)} \rightarrow 2 \text{NO}_2\text{(g)}$

The following data were obtained from three experiments using the method of initial rates:

| | Initial [NO] mol L ⁻¹ | Initial [O ₂] mol L ⁻¹ | Initial rate NO mol L ⁻¹ s ⁻¹ |
|--------------|-------------------------------------|--|--|
| Experiment 1 | 0.010 | 0.010 | 2.5×10^{-5} |
| Experiment 2 | 0.020 | 0.010 | 1.0×10^{-4} |
| Experiment 3 | 0.010 | 0.020 | 5.0×10^{-5} |

- Determine the order of the reaction for each reactant. **NO 2nd O₂ 1st**
- Write the rate equation for the reaction. **rate = k [NO]² [O₂]**
- Calculate the rate constant. **k = 25 L² mol⁻² s⁻¹**
- Calculate the rate (in mol L⁻¹s⁻¹) at the instant when [NO] = 0.015 mol L⁻¹ and [O₂] = 0.0050 mol L⁻¹. **rate = 2.8×10^{-5} mol L⁻¹ s⁻¹**
- At the instant when NO is reacting at the rate 1.0×10^{-4} mol L⁻¹s⁻¹, what is the rate at which O₂ is reacting and NO₂ is forming?
NO: -5.0×10^{-5} mol L⁻¹s⁻¹ ; NO₂: $+1.0 \times 10^{-4}$ mol L⁻¹s⁻¹

2. The reaction $2 \text{NO(g)} + 2 \text{H}_2\text{(g)} \rightarrow \text{N}_2\text{(g)} + 2 \text{H}_2\text{O(g)}$ was studied at 904 °C, and the data in the table were collected.

| | Initial [NO] mol L ⁻¹ | Initial [H ₂] mol L ⁻¹ | Initial rate N ₂ mol L ⁻¹ s ⁻¹ |
|--------------|-------------------------------------|--|--|
| Experiment 1 | 0.420 | 0.122 | 0.136 |
| Experiment 2 | 0.210 | 0.122 | 0.0339 |
| Experiment 3 | 0.210 | 0.244 | 0.0678 |
| Experiment 4 | 0.105 | 0.488 | 0.0339 |

- Determine the order of the reaction for each reactant. **NO 2nd H₂ 1st**
- Write the rate equation for the reaction. **rate = k [NO]² [H₂]**
- Calculate the rate constant at 904 °C. **k = 6.32 L² mol⁻² s⁻¹**
- Find the rate of appearance of N₂ at the instant when [NO] = 0.350 M and [H₂] = 0.205 M. **rate = 0.159 mol L⁻¹s⁻¹**

Rates Of Reaction 18 1 Practice Problems Worksheey

**Nancy L. Zucker, Katharine L.
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