Multiple Choices Quantum Numbers

A) energy level (shell) of the electron BI mass of the electron C) charge of the electron D) spin of the electron
ANS, A
 Which of the following is always true of all electrons that have the same principal quantum number.
A) they possess the same energy E) they have opposite spins C) they are in identical orbitals
D) all of the above E) none of the above
AMS/ III
1. An electron whose quantum numbers are $n=3,l=2$ and $ml=1$, is found in A) an s orbital B) a p orbital C) a d orbital D) an f orbital
ABISI C
1. Which of the following describes a 4p electron? A) (4,1,0,1/2)
ADVS: A.
1. If l=1, which of the following is not a possible value for m ₄ ? A) 0 80 1 C) -1 D) 2
ANS: D
1. The quantum number defines the shape of an orbital. A) **** **I0 magnetic C) principal C) secondary **** **** **** ***** ***** ***** ******
ANS. D
1. The azimuthal quantum number is 3 in orbitals. A) a B) p C) d D) f B) a
ANS: D
1. The principal quantum number of the first disubshell is A) 1 E) 2 C) 3 E) 4 E) 0
ANS, C

Quantum Chemistry Multiple Choice Questions

Arun Mittal

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