

1. Two groups of students (Group A and Group B) carried out a project* on the chemistry of some group 7 elements (the halogens) and their compounds.

- (a) In the first part of the project, the two groups had a sample of iodine monochloride (a corrosive brown liquid) prepared for them by their teacher using the following reaction.



The following data were recorded.

Mass of $\text{I}_2(\text{s})$	10.00 g
Mass of $\text{Cl}_2(\text{g})$	2.24 g
Mass of $\text{ICl}(\text{l})$ obtained	8.60 g

- (i) State the number of significant figures for the masses of $\text{I}_2(\text{s})$ and $\text{ICl}(\text{l})$. [1]

$\text{I}_2(\text{s})$:

$\text{ICl}(\text{l})$:

- (ii) The iodine used in the reaction was in excess. Determine the theoretical yield, in g, of $\text{ICl}(\text{l})$. [3]

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