

**52****ELECTRON ARRANGEMENT IN ATOMS****Section Review****Objectives**

- Describe how to write the electron configuration for an atom
- Explain why the actual electron configurations for some elements differ from those predicted by the Aufbau principle

**Vocabulary**

- electron configurations
- Aufbau principle
- Pauli exclusion principle
- Hund's rule

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The ways in which electrons are arranged around the nuclei of atoms are called 1. The 2 describes the sequence in which orbitals are filled. The various orbitals within a sublevel of a principle energy level are always of 3 energy. The 4 principle states that a maximum of only 5 electrons can occupy each orbital. To occupy the same orbital, two electrons must have 6 spins. Hund's rule states that the electrons pair up only after each orbital in a sublevel is occupied by 7. When using the shorthand method for showing the electron configuration of an atom, 8 are used to indicate the number of 9 occupying each sublevel.

Correct electron configurations can be obtained by using the Aufbau diagram for the elements up to and including vanadium. 10 and copper are exceptions to the Aufbau principle.

1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_
6. \_\_\_\_\_
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9. \_\_\_\_\_
10. \_\_\_\_\_

## Section 5 2 Electron Arrangement In Atoms Answer Key

**Donald Cotter**



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