



## Reinforcement

## Stability in Bonding

**Directions:** Each statement below contains a pair of terms or phrases in parentheses. Circle the term or phrase that makes each statement true.

- The properties of a compound are (the same as, different from) the properties of the elements that make up the compound.
- Na and Cl are chemical (symbols, formulas).
- NaCl and NaOH are chemical (symbols, formulas).
- In the formula  $H_2O$ , the number 2 is a (subscript, superscript).
- The number 2 in the formula  $H_2O$  tells you that each unit of this compound contains two (hydrogen, oxygen) atoms.
- If a symbol in a chemical formula does not have a subscript after it, a unit of that compound contains (no atoms, one atom) of that element.
- The total number of atoms in  $Fe_2O_3$  is (two, five, six).
- There are (three, seven, ten) different elements in  $H_2SO_4$ .
- An atom is chemically stable if its outer energy level (is filled with, contains no) electrons.
- For atoms of most noble gases and most other elements, the outer energy level is full when it has (3, 8) electrons.
- The noble gases do not readily form compounds because they (are, are not) chemically stable.
- A chemical bond is a (force, chemical) that holds atoms together in a compound.
- Chemical bonds form when atoms lose, gain, or (share, multiply) electrons.

**Directions:** Complete the table below by using the formula of each compound to identify the elements that each compound contains and the number of atoms of each of these elements in a unit of the compound. The first formula has been done for you.

Formula	Element 1	Element 2	Element 3
$H_2O$	2 hydrogen	1 oxygen	
14. NaOH			
15. NaCl			
16. $NH_3$			
17. $H_2SO_4$			
18. $SiO_2$			

# Section Stability In Bonding Answer Key

**Mr. Rohit Manglik**



## Section Stability In Bonding Answer Key:

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