

Section 5.3 Electron Configuration

In your textbook, read about ground-state electron configurations.

Use each of the terms below just once to complete the passage.

Aufbau principle	electron configuration	ground-state electron configuration	Hund's rule
lowest	Pauli exclusion principle	spins	stable

The arrangement of electrons in an atom is called the atom's

(1) _____. Electrons in an atom tend to assume the arrangement that gives the atom the (2) _____ possible energy. This arrangement of electrons is the most (3) _____ arrangement and is called the atom's (4) _____.

Three rules define how electrons can be arranged in an atom's orbitals. The

(5) _____ states that each electron occupies the lowest energy orbital available. The (6) _____ states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite (7) _____. (8) _____ states that single electrons with the same spin must occupy each equal-energy orbital before additional electrons with opposite spins occupy the same orbitals.

Complete the following table.

Element	Atomic Number	Orbitals					Electron Configuration
		1s	2s	2p _x	2p _y	2p _z	
9. Helium							1s ²
10.	7						
11. Neon		↑↓	↑↓	↑↓	↑↓	↑↓	

Section 5 3 Electron Configuration Answer Key

Gary Wulfsberg



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Understanding General Chemistry Atef Korchef, 2022-03-07 *Understanding General Chemistry* details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning, the book features main objectives at the beginning of each chapter, get smart sections, and check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs, exercises, and problems with the corresponding answers and explanations are readily available. Additional features include Alerts to students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self evaluate their skill level. Includes powerful methods, easy steps, simple and accurate interpretations, and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry, and physical chemistry. This introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students.

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