

Chapter 9 Carbon Chemistry

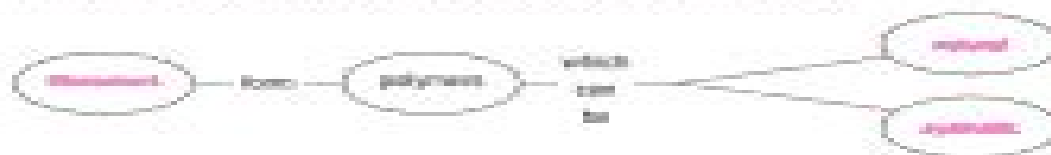
Section 9.3 Polymers

(pages 276–280)

This section explains how polymers form. It also discusses examples of synthetic and natural polymers.

Reading Strategy (page 276)

Identifying Main Ideas As you read, complete the concept map to summarize two main ideas about polymers. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.



- Describe a polymer. A polymer is a large molecule that forms when many smaller molecules are linked together by covalent bonds.
- The smaller molecules that join together to form a polymer are called monomers.
- Is the following sentence true or false? More than one type of monomer can be present in some polymers. True
- Name the two general classifications of polymers.
 - natural
 - synthetic

Synthetic Polymers (page 276)

- Name three polymers that can be synthesized.
 - Polyethylene
 - Nylon
 - Rubber
- Is the following sentence true or false? The more carbon atoms there are in a polyethylene chain, the harder the polymer is. True

Natural Polymers (pages 276–280)

- Name four types of polymers that are produced in plant and animal cells.
 - Starches
 - Complex proteins
 - Cellulose
 - Proteins
- Circle the letter of the molecular formula of a simple sugar.
 - CH_2O
 - $\text{C}_6\text{H}_{12}\text{O}_6$
 - $\text{C}_{12}\text{H}_{22}\text{O}_{11}$
 - $\text{C}_{10}\text{H}_{16}\text{O}_4$

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