

## Practice Problems

Do your best to solve all the following problems.

### SECTION 3.1 THE MOLE: A MEASUREMENT OF MATTER

- Calculate the molar mass of each of the following compounds:
  - phosphorus pentachloride ( $\text{PCl}_5$ )
  - carbon tetrachloride ( $\text{CCl}_4$ )
- Calculate the mass percent of each element in the compound:
  - $\text{KNO}_3$
  - $\text{Fe}_2\text{SO}_4$
- Calculate the number of moles of each element in 1.00 g of:
  - $\text{KNO}_3$
  - $\text{Fe}_2\text{SO}_4$
- Calculate the mass of 1.00 mole of each of the following:
  - hydrogen chloride gas ( $\text{HCl}$ )
  - hydrogen sulfide gas ( $\text{H}_2\text{S}$ )
  - hydrogen bromide gas ( $\text{HBr}$ )
  - hydrogen cyanide gas ( $\text{HCN}$ )

### SECTION 3.2 MOLE—MASS AND MOLE—VOLUME RELATIONSHIPS

- Calculate the mass of each of the following compounds:
 

a. $\text{H}_2\text{O}$	b. $\text{H}_2\text{SO}_4$	c. $\text{H}_2\text{SO}_3$	d. $\text{H}_2\text{SO}_4 \cdot \text{H}_2\text{O}$
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- Calculate the mass in grams of each of the following:
 

a. 0.00100 mole of $\text{H}_2\text{O}$	b. 0.00100 mole of $\text{H}_2\text{SO}_4$
c. 0.00100 mole of $\text{H}_2\text{SO}_3$	d. 0.00100 mole of $\text{H}_2\text{SO}_4 \cdot \text{H}_2\text{O}$
e. 0.00100 mole of $\text{H}_2\text{SO}_4$	f. 0.00100 mole of $\text{H}_2\text{SO}_3$
- Calculate the mass in grams of each of the following:
  - 0.00100 mole of  $\text{H}_2\text{O}$
  - 0.00100 mole of  $\text{H}_2\text{SO}_4$
  - 0.00100 mole of  $\text{H}_2\text{SO}_3$
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  - 0.00100 mole of  $\text{H}_2\text{SO}_3$
  - 0.00100 mole of  $\text{H}_2\text{SO}_4 \cdot \text{H}_2\text{O}$
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  - 0.00100 mole of  $\text{H}_2\text{SO}_3$
  - 0.00100 mole of  $\text{H}_2\text{SO}_4 \cdot \text{H}_2\text{O}$

# Section Chemical Quantities Practice Problems Answers

**Patrick E. McMahon, Rosemary  
McMahon, Bohdan Khomtchouk**



**Section Chemical Quantities Practice Problems Answers:**

## Unveiling the Magic of Words: A Report on "**Section Chemical Quantities Practice Problems Answers**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Section Chemical Quantities Practice Problems Answers**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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