

CHAPTER 2 THE COMPONENTS OF MATTER

END-OF-CHAPTER PROBLEMS

- 2.1 **Plan:** Refer to the definitions of an element and a compound.
Solution:
Unlike compounds, elements cannot be broken down by chemical changes into simpler materials. Compounds contain different types of atoms; there is only one type of atom in an element.
- 2.2 **Plan:** Refer to the definitions of a compound and a mixture.
Solution:
1) A compound has constant composition but a mixture has variable composition. 2) A compound has distinctly different properties than its component elements; the components in a mixture retain their individual properties.
- 2.3 **Plan:** Recall that a substance has a fixed composition.
Solution:
a) The fixed mass ratio means it has constant composition, thus, it is a **pure substance** (compound).
b) All the atoms are identical, thus, it is a **pure substance** (element).
c) The composition can vary, thus, this is an **impure substance** (a mixture).
d) The specific arrangement of different atoms means it has constant composition, thus, it is a **pure substance** (compound).
- 2.4 **Plan:** Remember that an element contains only one kind of atom while a compound contains at least two different elements (two kinds of atoms) in a fixed ratio. A mixture contains at least two different substances in a composition that can vary.
Solution:
a) The presence of more than one element (calcium and chlorine) makes this pure substance a **compound**.
b) There are only atoms from one element, sulfur, so this pure substance is an **element**.
c) This is a combination of two compounds and has a varying composition, so this is a **mixture**.
d) The presence of more than one type of atom means it cannot be an element. The specific, not variable, arrangement means it is a **compound**.
- 2.5 **Plan:** Recall that an element contains only one kind of atom; the atoms in an element may occur as molecules. A compound contains two kinds of atoms (different elements).
Solution:
a) This scene has 3 atoms of an element, 2 molecules of one compound (with one atom each of two different elements), and 2 molecules of a second compound (with 2 atoms of one element and one atom of a second element).
b) This scene has 2 atoms of one element, 2 molecules of a diatomic element, and 2 molecules of a compound (with one atom each of two different elements).
c) This scene has 2 molecules composed of 3 atoms of one element and 3 diatomic molecules of the same element.
- 2.6 **Plan:** Restate the three laws in your own words.

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Puja Mehta



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