

Solutions, Acid/Base & Solubility Test Review

Know the definitions for the following:

Solvent – liquid (usually water) that dissolves the solute to make the solution
Solute – solid (usually) that is added to solvent and dissolves to make solution
Solution – homogeneous mixture of solvent and solute
Concentration – a measure of how much solute per volume of solvent
Molarity – concentration; moles of solute/liters of solvent
Arrhenius Acid and Base – acid donates H^+ and base donates OH^-
Bronsted Lowry Acid and Base – acid donates H^+ and a base accepts H^+
Neutralization – acid + base \rightarrow salt + water
Titration – process used to determine concentration of an acid using neutralization

Solutions/Solubility:

What factors will increase the rate of solubility? Increasing temperature; breaking into pieces; stirring

Calculating Concentration (M = moles/liters)

What is the molarity of 50 grams of NaCl in a 100ml of water?

Convert grams \rightarrow moles: 50 grams NaCl / 58.5 grams (from periodic table) = .85 moles
Divide moles by volume \rightarrow .85 moles / 0.1 L = 8.5 M

Dilutions ($M_1V_1 = M_2V_2$)

If I have 500 mL of a 0.5M NaCl solution, what will the concentration be if the final volume of the solution is 600 mL? $(500\text{ml}) \cdot (0.5\text{M}) = (600\text{ml})(? \text{ M}) = 0.42 \text{ M}$

How many liters of 0.5M HCl stock solution would you need to prepare 2 liters of 0.25 M HCl solution? $(2 \text{ L}) \cdot (0.25 \text{ M}) = (? \text{ L}) \cdot (0.5 \text{ M}) = 1 \text{ M}$

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