

CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS

FOLLOW-UP PROBLEMS

17.1A **Plan:** First, balance the equations and then write the reaction quotient. Products appear in the numerator of the reaction quotient and reactants appear in the denominator; coefficients in the balanced reaction become exponents.

Solution:



$$\text{Reaction quotient: } Q_c = \frac{[\text{NO}]^4 [\text{H}_2\text{O}]^6}{[\text{NH}_3]^4 [\text{O}_2]^5}$$



$$\text{Reaction quotient: } Q_c = \frac{[\text{NH}_3]^2}{[\text{NO}_2]^2 [\text{H}_2]^7}$$



$$\text{Reaction quotient: } Q_c = [\text{O}_2]^3$$

17.1B **Plan:** First, balance the equations and then write the reaction quotient. Products appear in the numerator of the reaction quotient and reactants appear in the denominator; coefficients in the balanced reaction become exponents.

Solution:



$$\text{Reaction quotient: } Q_c = \frac{[\text{CO}]^2 [\text{H}_2]^2}{[\text{CH}_4] [\text{CO}_2]}$$



$$\text{Reaction quotient: } Q_c = \frac{[\text{H}_2\text{O}]^2}{[\text{H}_2\text{S}]^2 [\text{SO}_2]}$$



$$\text{Reaction quotient: } Q_c = \frac{[\text{NaCN}]}{[\text{HCN}] [\text{NaOH}]}$$

17.2A **Plan:** Add the individual steps to find the overall equation, canceling substances that appear on both sides of the equation. Write the reaction quotient for each step and the overall equation. Multiply the reaction quotients for each step and cancel terms to obtain the overall reaction quotient.

Solution:



Canceling the reactants leaves the overall equation as $\text{Br}_2(g) + \text{H}_2(g) \rightleftharpoons 2\text{HBr}(g)$.

Write the reaction quotients for each step:

$$Q_{c1} = \frac{[\text{Br}]^2}{[\text{Br}_2]} \quad Q_{c2} = \frac{[\text{HBr}][\text{H}]}{[\text{Br}][\text{H}_2]} \quad Q_{c3} = \frac{[\text{HBr}]}{[\text{H}][\text{Br}]}$$

and for the overall equation:

$$Q_c = \frac{[\text{HBr}]^2}{[\text{H}_2][\text{Br}_2]}$$

Solutions Manual Chemical Equilibrium

David Smith



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Solutions Manual to Accompany Elements of Physical Chemistry David Smith, 2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions *Solutions Manual - Chemistry* Charles E. Mortimer, 1986 *Solutions Manual for Principles of Physical Chemistry, 3rd Edition* Hans Kuhn, David H. Waldeck, Horst-Dieter Försterling, 2024-10-29 This is a Solutions Manual to Accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum mechanics atomic and molecular structure atomic and molecular spectroscopy statistical thermodynamics classical thermodynamics solutions and equilibria electrochemistry kinetics and reaction dynamics macromolecules and organized molecular assemblies Importantly the book develops and applies these principles to supramolecular assemblies and supramolecular machines with many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text Solutions Manual for Principles of Physical Chemistry, 3rd Edition, Solutions Manual Hans Kuhn, David H. Waldeck, Horst-Dieter Försterling, 2024-10-25 This is a Solutions Manual to Accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum

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