

## WS 2: Writing Balanced Chemical Equations

Step 1: Circle the reactants and underline the products in the problems below.

Step 2: Write the unbalanced equations; show all states of matter.

Step 3: Balance the equation

1. Fluorine gas reacts with solid sodium chloride to form solid sodium fluoride and chlorine gas.



2. Nitrous oxide gas (systematic name: dinitrogen monoxide) and water vapor are produced by carefully heating solid ammonium nitrate.



3. At high temperatures and in the presence of a catalyst, nobles gases can be made to form stable compounds. Such an example is xenon gas which combines with fluorine gas to produce solid xenon tetrafluoride.



4. Liquid phosphorus trichloride reacts with water to form phosphorous acid and hydrochloric acid in solution.



5. Solid manganese(II) peroxide reacts with aqueous hydrochloric acid to form manganese(II) chloride, chlorine gas, and water.



6. When aqueous solutions of potassium chromate and barium chloride are mixed, barium chromate settles out of solution as a solid leaving potassium chloride in solution.



# Writing Balanced Chemical Equations Answers

**D Kirk**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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