

WRITING CHEMICAL EQUATIONS

Name _____	Section _____	Date _____
Writing Equations		
Directions: Write and balance each chemical equation. You will need to write the products. Include phase labels with each reactant and product. You may need to refer to a table of solubility rules.		
1) Aqueous sulfuric acid reacts with aqueous potassium hydroxide		
2) Aqueous potassium carbonate reacts with aqueous hydrochloric acid		
3) Hydrogen and oxygen gases combine to form water vapor		

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1) Aqueous sulfuric acid reacts with aqueous potassium hydroxide		
2) Aqueous potassium carbonate reacts with aqueous hydrochloric acid		
3) Hydrogen and oxygen gases combine to form water vapor		
4) Methane gas (carbon tetrahydride) reacts with chlorine gas to yield chloroform (CH ₃ Cl) gas and hydrochloric acid gas		
5) Aqueous ammonium sulfide reacts with aqueous magnesium sulfate to yield solid magnesium sulfide and aqueous ammonium sulfate		
6) Mercury (II) nitrate reacts with aqueous iron(III) iodide to form aqueous iron (III) nitrate and solid mercury (II) iodide		
7) Aqueous strontium hydroxide reacts with aqueous iron (III) nitrate to produce solid iron hydroxide and aqueous strontium nitrate		
8) Solid zinc carbonate reacts with aqueous potassium chloride. An aqueous solution of zinc chloride and potassium carbonate is formed.		
9) Aqueous ammonium chromate reacts with aqueous barium sulfide react to form aqueous ammonium sulfide and solid barium chromate.		
10) Aqueous lithium oxalate reacts with aqueous barium sulfate to form aqueous lithium sulfate and solid barium oxalate		

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A CHEMISTRY WORKSHEET

10 PROBLEMS

WORKED ANSWER KEY INCLUDED

Writing Chemical Equations And Answers

Dr. Sunita Bhagiya, ,Er. Meera Goyal



Writing Chemical Equations And Answers:

Chemistry Therald Moeller, 2012-12-02 Chemistry with Inorganic Qualitative Analysis is a textbook that describes the application of the principles of equilibrium represented in qualitative analysis and the properties of ions arising from the reactions of the analysis This book reviews the chemistry of inorganic substances as the science of matter the units of measure used atoms atomic structure thermochemistry nuclear chemistry molecules and ions in action This text also describes the chemical bonds the representative elements the changes of state water and the hydrosphere which also covers water pollution and water purification Water purification occurs in nature through the usual water cycle and by the action of microorganisms The air flushes dissolved gases and volatile pollutants when water seeps through the soil it filters solids as they settle in the bottom of placid lakes Microorganisms break down large organic molecules containing mostly carbon hydrogen nitrogen oxygen sulfur or phosphorus into harmless molecules and ions This text notes that natural purification occurs if the level of contaminants is not so excessive This textbook is suitable for both chemistry teachers and students

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