Stoichiometric Calculations

Mole to Mole Calculations

The octane present in gasoline burns according to the following equation.

$$2C_8H_{18} + 25O_2 \rightarrow 16CO_2 + 18H_2O$$

- (a) How many moles of O₂ are needed to react fully with 7 moles of octane (C₂H₁₈)?
- (b) How many moles of CO₂ can form from 5 mole of octane?
- (c) If the reaction is to be used to synthesize 8 moles of CO₂, how many moles of oxygen are needed?
- (d) How many moles of carbon dioxide are formed if 17 moles of oxygen are reacted?
- Butane combines with oxygen in a combustion reaction to produce carbon dioxide and water.
 - (a) Write the balanced equation for this reaction.
 - (b) If 37 moles of water are produced, how many moles of butane were reacted?
 - (c) 29 moles of oxygen are reacted with the butane, how many moles of carbon dioxide are produced?
 - (d) How many moles of butane are required to react with 17.5 moles of oxygen?
- Consider the following reaction:

$$3Cu + 8HNO_3 \rightarrow 3Cu(NO_3)_2 + 2NO + 4H_2O$$

- (a) How many moles of NO are produced by the reaction of 4.0 moles of copper with HNO₃?
- (b) How many moles of HNO₃ are required to react completely with 5.0 mol of copper?
- (c) What mass of NO is produced when 6.35 g of copper is reacted with HNO₃?

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