

$$M_{\text{AgNO}_3} = \frac{n_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{25.513 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{1.569 \frac{\text{mol}}{\text{L}}}$$

4. Confirm that the molarity is the same molarity that you calculated in step 3.

Yes, the molarity is the same as the molarity calculated in step 3.

### Part B: Unknown solutions

Using a method similar to the one used in part A, create and carry out a procedure for determining the molarities of the unknown  $\text{AgNO}_3$  solution and the unknown  $\text{NaCl}$  solution. Round your data and show your calculations.

#### Unknown $\text{AgNO}_3$

Upon the addition of solid  $\text{NaCl}$ , solid  $\text{AgCl}$  was formed.

$\text{AgCl}$  formed (g) = 2.32019 g

$$\text{mass}(\text{AgNO}_3) = \frac{M_r(\text{AgNO}_3)}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{149.87}{143.32} \frac{\text{g}}{\text{mol}} \times 2.32019 \text{ g} = \boxed{2.3500 \text{ g}}$$

$$M_{\text{AgNO}_3} = \frac{n_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{2.3500 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.16189 \frac{\text{mol}}{\text{L}}}$$

#### Unknown $\text{NaCl}$

Upon the addition of solid  $\text{AgNO}_3$ , solid  $\text{AgCl}$  was formed.

$\text{AgCl}$  formed (g) = 10.7907 g

$$\text{mass}(\text{NaCl}) = \frac{M_r(\text{NaCl})}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{58.443}{143.32} \frac{\text{g}}{\text{mol}} \times 10.7907 \text{ g} = \boxed{4.4002 \text{ g}}$$

$$M_{\text{NaCl}} = \frac{n_{\text{NaCl}}}{V_{\text{NaCl}}} = \frac{4.4002 \text{ g} + 58.443 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.75294 \frac{\text{mol}}{\text{L}}}$$

### Conclusion questions

1. If you have a company that is using this method to create  $\text{AgCl}$ , knowing that  $\text{AgNO}_3$  is an expensive chemical, what do you think would be a better limiting reagent to use,  $\text{AgNO}_3$  or  $\text{NaCl}$ ? Why?

# Stoichiometry Limiting Reagent Pre Lab Answers

**Michal Rosen-Zvi**



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