Chemistry

Name

is datters

Stoichiometry Reflect #2

Learning Target... I can use stoichiometry to calculate amounts

- 1a. Balance this equation $A = \frac{1}{2} \left(\frac{1}{N} + \frac{3}{2} \right) \right) \right) \right) \right) \right) \right) \right) \right]$
- 1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 q of it.

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

1d. Determine how many grams of ammonium nitride produced in this reaction.

Using the following chemical reaction: 2 Hg + Cu(NO₃)₂ → 2 HgNO₃ + Cu ↑ A and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

3, Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!) $\eta_{0} = \eta_{0} = 1.75$

2 Al + 3 Cucle > 3 Cu + 2 AlCl3 B

=[91.4%]

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ACT

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