

Stoichiometry Problem Solving Steps

STOICHIOMETRY STEPS

1. Write a balanced equation
 - a. This gives you the ratio of atoms and molecules that react (coefficients)
2. Write what I know and what I want to know ABOVE the equation
 - a. This tells you how to solve the problem
3. Go down to moles
 - a. USE TRIANGLES TO GET UNITS OF UNKNOWN
4. Go over to moles
 - a. USE TRIANGLES TO GET UNITS OF UNKNOWN
5. Go up to the answer
 - a. USE TRIANGLES TO GET UNITS OF UNKNOWN

EXAMPLE

I have 4.5 grams of aluminum. How many grams of sulfur will I need to react and produce maximum product?

First, write a BALANCED EQUATION.



Second, write what you know and what to know ABOVE the equation.



Third, go DOWN to moles using TRIANGLES.

$$\frac{4.5 \text{ g}}{2 \text{ Al (s)}} \times \frac{1 \text{ mol Al}}{26.98 \text{ g/mol Al}} = 0.147 \text{ mol Al}$$



Fourth, go OVER to moles using MOLE RATIO.

$$0.147 \text{ mol Al} \times \frac{3 \text{ mol S}}{2 \text{ mol Al}} = 0.221 \text{ mol S}$$



Fifth, go UP to the answer using TRIANGLES.

$$0.221 \text{ mol S} \times \frac{32.07 \text{ g S}}{1 \text{ mol S}} = 7.09 \text{ g S}$$



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R.D Tomlinson,A.E Hill,R.D Pilkington



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