

Name: \_\_\_\_\_ Period: \_\_\_\_\_

1. The common pain medicine, Advil, contains the active ingredient ibuprofen that has a molar mass of 206 g/mol. Ibuprofen contains 75.73% C 8.74% H the remaining mass being oxygen. What are the empirical and molecular formulae for ibuprofen?

Assume 100g 75.73g C 8.74g H 15.53g O

$$75.73g C \cdot \frac{1 \text{ mol C}}{12.01g C} = 6.31 \text{ mol C}$$

$$8.74g H \cdot \frac{1 \text{ mol H}}{1.008g H} = 8.67 \text{ mol H}$$

$$15.53g O \cdot \frac{1 \text{ mol O}}{16g O} = 0.971 \text{ mol O}$$

$$\frac{6.31 \text{ mol C}}{0.971 \text{ mol}} =$$

$$\frac{8.67 \text{ mol H}}{0.971 \text{ mol}} =$$

$$\frac{0.971 \text{ mol O}}{0.971 \text{ mol}} =$$

6.5 mol C

8.93 mol H

1 mol O

x2 13 mol C

18 mol H

2 mol O

$C_{13}H_{18}O_2$  ← empirical + molecular formula

# Stoichiometry Review Answer Key

**Brendan G. Carr**



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