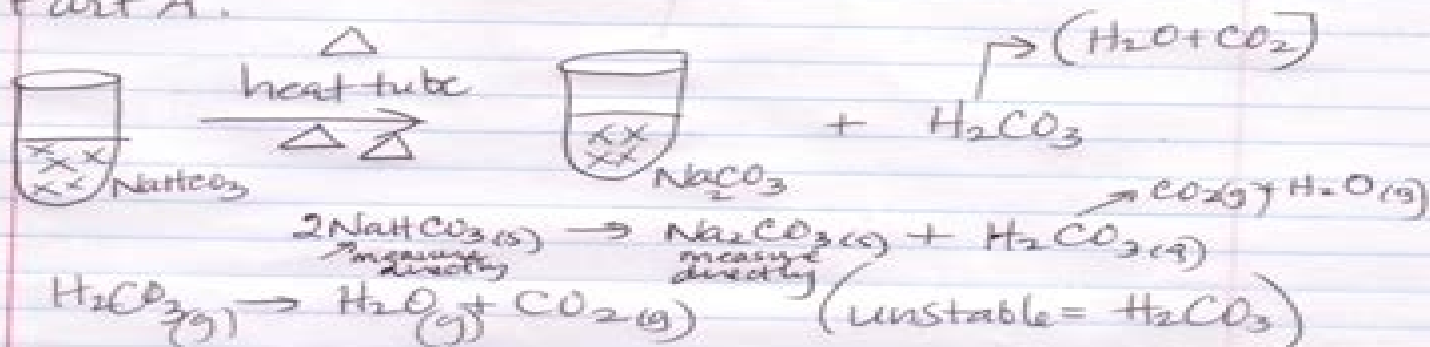


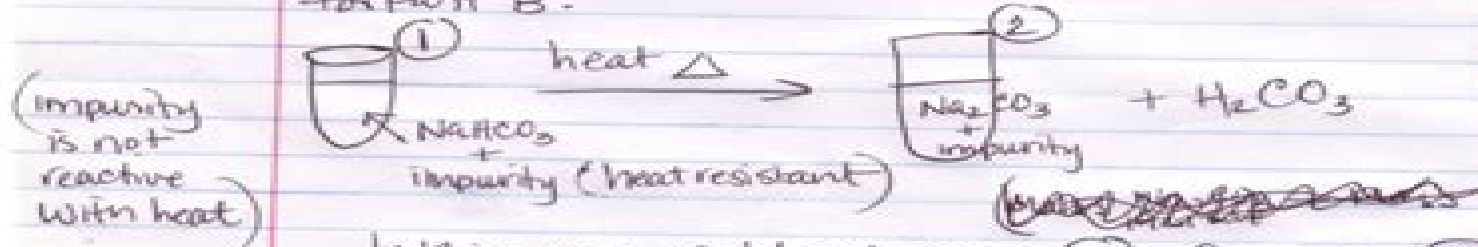
Conceptual: Baking Soda Lab

Part A.



For Part A, ~~they don't ask you to calculate this~~
~~mass lost = [Mass of Test tube + NaHCO_3] - [Mass of Test tube + Na_2CO_3]~~
 (or mass H_2CO_3)
 you can measure product in part A without this because it is pure material

For Part B.



In this case, subtract mass (2) from mass (1) to get mass H_2CO_3 ($\text{CO}_2 + \text{H}_2\text{O}$).
 $(\text{NaHCO}_3 + \text{test tube}) - (\text{Na}_2\text{CO}_3 + \text{test tube}) = \text{mass H}_2\text{CO}_3 \text{ lost}$
 Then, work backwards (see 2nd example calculation) to get mass of pure NaHCO_3 in measured sample.

$$\frac{\text{Mass pure NaHCO}_3}{\text{Mass impure NaHCO}_3 + \text{impurity}} \times 100 = \% \text{ composition}$$

Stoichiometry Lab Vinegar And Baking Soda Answers

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