

STOICHIOMETRY: MASS-MASS PROBLEMS

Name _____



How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose?

$$25\text{g KClO}_3 \left(\frac{1\text{ mol}}{122.5\text{g}} \right) \left(\frac{2\text{ mol KCl}}{2\text{ mol KClO}_3} \right) \left(\frac{74.5\text{g KCl}}{1\text{ mol KCl}} \right) = 13.2\text{g KCl}$$



How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

$$50\text{g N}_2 \left(\frac{1\text{ mol N}_2}{28\text{g N}_2} \right) \left(\frac{3\text{ mol H}_2}{1\text{ mol N}_2} \right) \left(\frac{2\text{g H}_2}{1\text{ mol H}_2} \right) = 10.7\text{g H}_2$$

3. How many grams of ammonia are produced in the reaction in Problem 2?

$$50\text{g N}_2 \quad 50\text{g} + 10.7\text{g} = \boxed{60.7\text{g NH}_3}$$



How many grams of silver chloride are produced from 5.0 g of silver nitrate reacting with an excess of barium chloride?

$$5.0\text{g AgNO}_3 \left(\frac{1\text{ mol}}{170\text{g}} \right) \left(\frac{2\text{ mol AgCl}}{2\text{ mol AgNO}_3} \right) \left(\frac{143\text{g}}{1\text{ mol AgCl}} \right) = 4.21\text{g AgCl}$$

5. How much barium chloride is necessary to react with the silver nitrate in Problem 4?

$$5.0\text{g AgNO}_3 \left(\frac{1\text{ mol}}{170\text{g AgNO}_3} \right) \left(\frac{1\text{ mol BaCl}_2}{2\text{ mol AgNO}_3} \right) \left(\frac{207\text{g}}{1\text{ mol BaCl}_2} \right) = \boxed{5.0\text{g BaCl}_2}$$

9-7

Stoichiometry Mixed Problems Answer Key

Terry C. Jones



Stoichiometry Mixed Problems Answer Key:

This book delves into Stoichiometry Mixed Problems Answer Key. Stoichiometry Mixed Problems Answer Key is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Stoichiometry Mixed Problems Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Stoichiometry Mixed Problems Answer Key
- Chapter 2: Essential Elements of Stoichiometry Mixed Problems Answer Key
- Chapter 3: Stoichiometry Mixed Problems Answer Key in Everyday Life
- Chapter 4: Stoichiometry Mixed Problems Answer Key in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Stoichiometry Mixed Problems Answer Key. The first chapter will explore what Stoichiometry Mixed Problems Answer Key is, why Stoichiometry Mixed Problems Answer Key is vital, and how to effectively learn about Stoichiometry Mixed Problems Answer Key.

3. In chapter 2, this book will delve into the foundational concepts of Stoichiometry Mixed Problems Answer Key. This chapter will elucidate the essential principles that need to be understood to grasp Stoichiometry Mixed Problems Answer Key in its entirety.

4. In chapter 3, this book will examine the practical applications of Stoichiometry Mixed Problems Answer Key in daily life. This chapter will showcase real-world examples of how Stoichiometry Mixed Problems Answer Key can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Stoichiometry Mixed Problems Answer Key in specific contexts. The fourth chapter will explore how Stoichiometry Mixed Problems Answer Key is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Stoichiometry Mixed Problems Answer Key. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Stoichiometry Mixed Problems Answer Key.

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