

# Stoichiometry "End of Chapter" IN CLASS Practice Assessment

## Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?  
Coefficients are necessary for establishing a molar ratio to convert between moles of different substances.

**Skill 1** - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

**Skill 2** - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[ \frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[ \frac{2.8\text{g}}{4.0\text{g}} \right] \times 100 = 70\%$$

**Skill 3** - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol} \\ 250\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} = 6.857\text{mol HCl} \quad \text{2mol in rxn} = 3.43\text{mol}$$

$$1.55\text{mol CaCO}_3 \times \left( \frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left( \frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[ \frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[ \frac{142.0\text{g}}{172.02\text{g}} \right] \times 100 = 82.55\%$$

# Stoichiometry Form A Test Answer Key

**Xiang Xie**



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## **Table of Contents Stoichiometry Form A Test Answer Key**

1. Understanding the eBook Stoichiometry Form A Test Answer Key
  - The Rise of Digital Reading Stoichiometry Form A Test Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Form A Test Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Stoichiometry Form A Test Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Form A Test Answer Key
  - Personalized Recommendations
  - Stoichiometry Form A Test Answer Key User Reviews and Ratings
  - Stoichiometry Form A Test Answer Key and Bestseller Lists
5. Accessing Stoichiometry Form A Test Answer Key Free and Paid eBooks
  - Stoichiometry Form A Test Answer Key Public Domain eBooks
  - Stoichiometry Form A Test Answer Key eBook Subscription Services
  - Stoichiometry Form A Test Answer Key Budget-Friendly Options

6. Navigating Stoichiometry Form A Test Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Stoichiometry Form A Test Answer Key Compatibility with Devices
  - Stoichiometry Form A Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Stoichiometry Form A Test Answer Key
  - Highlighting and Note-Taking Stoichiometry Form A Test Answer Key
  - Interactive Elements Stoichiometry Form A Test Answer Key
8. Staying Engaged with Stoichiometry Form A Test Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Stoichiometry Form A Test Answer Key
9. Balancing eBooks and Physical Books Stoichiometry Form A Test Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Stoichiometry Form A Test Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Form A Test Answer Key
  - Setting Reading Goals Stoichiometry Form A Test Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Form A Test Answer Key
  - Fact-Checking eBook Content of Stoichiometry Form A Test Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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