

## Stoichiometry Practice 2.1 - Answer Key

**Objectives:** - Use stoichiometry to calculate the mass of any substance involved in a chemical reaction

**Directions:** Answer the following stoichiometry problems and show your work.

1. If a piece of magnesium with a mass of 2.76 g is added to a solution of hydrochloric acid (HCl), what mass of  $H_2$  would be produced? **[0.229 g]**



$$2.76 \text{ g Mg} = \left( \frac{1 \text{ mol Mg}}{24.305 \text{ g Mg}} \right) \left( \frac{1 \text{ mol } H_2}{1 \text{ mol Mg}} \right) \left( \frac{2.02 \text{ g } H_2}{1 \text{ mol } H_2} \right) = \boxed{0.229 \text{ g } H_2}$$

2. What mass of barium chloride is needed to react completely with 46.8 g of sodium phosphate?



$$46.8 \text{ g } Na_3PO_4 = \left( \frac{1 \text{ mol } Na_3PO_4}{163.94 \text{ g } Na_3PO_4} \right) \left( \frac{3 \text{ mol } BaCl_2}{2 \text{ mol } Na_3PO_4} \right) \left( \frac{208.23 \text{ g } BaCl_2}{1 \text{ mol } BaCl_2} \right) = \boxed{89.1 \text{ g } BaCl_2}$$

3. In the reaction, where magnesium oxide reacts with water to produce magnesium hydroxide. What mass of magnesium oxide is needed to produce 264 g of magnesium hydroxide?



$$264 \text{ g } Mg(OH)_2 = \left( \frac{1 \text{ mol } Mg(OH)_2}{58.32 \text{ g } Mg(OH)_2} \right) \left( \frac{1 \text{ mol } MgO}{1 \text{ mol } Mg(OH)_2} \right) \left( \frac{40.30 \text{ g } MgO}{1 \text{ mol } MgO} \right) = \boxed{18.2 \text{ g } MgO}$$

4. Methane ( $CH_4$ ) burns in oxygen to produce carbon dioxide and water. What mass of oxygen would be necessary for the reaction if you had 15.0 g of methane?



$$15.0 \text{ g } CH_4 = \left( \frac{1 \text{ mol } CH_4}{16.04 \text{ g } CH_4} \right) \left( \frac{2 \text{ mol } O_2}{1 \text{ mol } CH_4} \right) \left( \frac{32.00 \text{ g } O_2}{1 \text{ mol } O_2} \right) = \boxed{59.8 \text{ g } O_2}$$

# Stoichiometry Mass Mass Problems Worksheet Answers

**NARAYAN CHANGDER**



## Stoichiometry Mass Mass Problems Worksheet Answers:

**General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material Stoichiometry Unit Project Luann Marie Decker, 1998 *Holt Chemistry* Holt Rinehart & Winston, 2003-01-24 *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **A Concrete Stoichiometry Unit for High School Chemistry** Jennifer Louise Pakkala, 2006 Holt Chemistry Ralph Thomas Myers, 2004 *Chemistry Study Guide* Gregory Gebhart, 2017-09-21 Solutions to stoichiometry and conservation of mass problems Laboratory exercises Bite-Sized Chemistry Calculations John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume Volume Stoichiometry 13 Percent Yield Answers to Practice Problems Appendices *STOICHIOMETRY* NARAYAN CHANGDER, 2024-04-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will

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**A Stoichiometry Unit** David Callaghan,2004 *You Can Do Chemistry* Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

**Chemistry Games** Gregory Gebhardt,2012-10-01 This book is a science education text It is a collection of Chemistry games that teaches the periodic table stoichiometry chemical reactions and the law of conservation of mass

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