

Stoichiometry Practice Worksheet

Solve the following stoichiometry grams-grams problems:

- 1) Using the following equation: $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

- 2) Using the following equation: $\text{Pb}(\text{SO}_4)_2 + 4 \text{ LiNO}_3 \rightarrow \text{Pb}(\text{NO}_3)_4 + 2 \text{ Li}_2\text{SO}_4$

How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

Limiting Reagent Worksheet

For the following reactions, find the following:

- a) Which of the reagents is the limiting reagent?
b) What is the maximum amount of each product that can be formed?
c) How much of the other reagent is left over after the reaction is complete?

- 3) Consider the following reaction:



Answer the questions above (a,b,c), assuming we started with 30 grams of ammonium nitrate and 50 grams of sodium phosphate.

- 4) Consider the following reaction:



Answer the questions above (a,b,c) assuming we start with 100 grams of calcium carbonate and 45 grams of iron (III) phosphate.

5) When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

- a) Write the balanced equation for the reaction given above:
b) If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?
c) What is the limiting reagent for the reaction in #2? _____
d) How much of the nonlimiting reagent is left over in this reaction?

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Arshad Iqbal



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Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner, 2007 **A**

Stoichiometry Unit David Callaghan, 2004 You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the

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