

Unit 8 – Stoichiometry Worksheet 1: Mole relationships

For each of the problems below, use the five steps discussed in class. Answer on a separate piece of paper.

1. Write the balanced chemical equation
2. Identify what is given (with units) and what you want to find (with units) in a BCA table
3. Use the mole ratio from the balanced equation to calculate remaining values in the Table (dimensional analysis style). **Show your work beneath the BCA table.**
4. Complete the BCA Table using calculation results
5. Find the desired quantity in your work. Write and box in that answer with unit, substance.

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8.0 moles of hydrogen sulfide? (The BCA table is done as an example. Make/copy your own on your own piece of paper).

Equation: $\text{___ H}_{2}\text{S}_{(g)} + \text{___ O}_{2(g)} \rightarrow \text{___ SO}_{2(g)} + \text{___ H}_2\text{O}_{(g)}$

Before:

FILL IN THE VALUES ON YOUR OWN PAPER

Change

(SHOW WORK UNDER THE TABLE)

After

2. Propane, C_3H_8 , burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?
3. Ammonia, NH_3 , for fertilizer is produced by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?
4. The poison gas phosgene, COCl_2 , reacts with water in the lungs to form hydrochloric acid and carbon dioxide. How many moles of hydrochloric acid would be formed by 0.835 moles of phosgene?
5. Nitric acid, HNO_3 , is used in the production of trinitrotoluene, (TNT), an explosive. The equation (not balanced) is



- a) How many moles of nitric acid are required to react with 4.5 moles of toluene? (make your own BCA table)
 - b) How many moles of TNT should be produced?
6. Iron metal and oxygen combine to form magnetite, the magnetic oxide of iron, Fe_3O_4 .
 - a) How many moles of iron can be converted to magnetite by 8.80 moles of pure oxygen? (make your own BCA table)
 - b) How many moles of the magnetic iron oxide would be produced?

Stoichiometry Unit 8 Mole Relationships Answers

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Table of Contents Stoichiometry Unit 8 Mole Relationships Answers

1. Understanding the eBook Stoichiometry Unit 8 Mole Relationships Answers
 - The Rise of Digital Reading Stoichiometry Unit 8 Mole Relationships Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Unit 8 Mole Relationships Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Unit 8 Mole Relationships Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Unit 8 Mole Relationships Answers
 - Personalized Recommendations
 - Stoichiometry Unit 8 Mole Relationships Answers User Reviews and Ratings
 - Stoichiometry Unit 8 Mole Relationships Answers and Bestseller Lists

5. Accessing Stoichiometry Unit 8 Mole Relationships Answers Free and Paid eBooks
 - Stoichiometry Unit 8 Mole Relationships Answers Public Domain eBooks
 - Stoichiometry Unit 8 Mole Relationships Answers eBook Subscription Services
 - Stoichiometry Unit 8 Mole Relationships Answers Budget-Friendly Options
6. Navigating Stoichiometry Unit 8 Mole Relationships Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Unit 8 Mole Relationships Answers Compatibility with Devices
 - Stoichiometry Unit 8 Mole Relationships Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Unit 8 Mole Relationships Answers
 - Highlighting and Note-Taking Stoichiometry Unit 8 Mole Relationships Answers
 - Interactive Elements Stoichiometry Unit 8 Mole Relationships Answers
8. Staying Engaged with Stoichiometry Unit 8 Mole Relationships Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Unit 8 Mole Relationships Answers
9. Balancing eBooks and Physical Books Stoichiometry Unit 8 Mole Relationships Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Unit 8 Mole Relationships Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Unit 8 Mole Relationships Answers
 - Setting Reading Goals Stoichiometry Unit 8 Mole Relationships Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Unit 8 Mole Relationships Answers
 - Fact-Checking eBook Content of Stoichiometry Unit 8 Mole Relationships Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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