

Balancing Equations Worksheet

- 1) $\text{H}_3\text{PO}_4 + \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$
- 2) $\text{K} + \text{B}_2\text{O}_3 \rightarrow \text{K}_2\text{O} + \text{B}$
- 3) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- 4) $\text{Na} + \text{NaNO}_3 \rightarrow \text{Na}_2\text{O} + \text{N}_2$
- 5) $\text{C} + \text{S}_8 \rightarrow \text{CS}_2$
- 6) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 7) $\text{N}_2 + \text{O}_2 \rightarrow \text{N}_2\text{O}_5$
- 8) $\text{H}_3\text{PO}_4 + \text{Mg}(\text{OH})_2 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 9) $\text{NaOH} + \text{H}_2\text{CO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$
- 10) $\text{KOH} + \text{HBr} \rightarrow \text{KBr} + \text{H}_2\text{O}$
- 11) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 12) $\text{Al}(\text{OH})_3 + \text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + \text{H}_2\text{O}$
- 13) $\text{Al} + \text{S}_8 \rightarrow \text{Al}_2\text{S}_3$
- 14) $\text{Cs} + \text{N}_2 \rightarrow \text{Cs}_3\text{N}$
- 15) $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
- 16) $\text{Rb} + \text{RbNO}_3 \rightarrow \text{Rb}_2\text{O} + \text{N}_2$
- 17) $\text{C}_6\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 18) $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 19) $\text{C}_{10}\text{H}_{22} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 20) $\text{Al}(\text{OH})_3 + \text{HBr} \rightarrow \text{AlBr}_3 + \text{H}_2\text{O}$
- 21) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 22) $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 23) $\text{Li} + \text{AlCl}_3 \rightarrow \text{LiCl} + \text{Al}$
- 24) $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 25) $\text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + \text{H}_2\text{O}$
- 26) $\text{Rb} + \text{P} \rightarrow \text{Rb}_3\text{P}$
- 27) $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 28) $\text{Al}(\text{OH})_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- 29) $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
- 30) $\text{Rb} + \text{S}_8 \rightarrow \text{Rb}_2\text{S}$
- 31) $\text{H}_3\text{PO}_4 + \text{Ca}(\text{OH})_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 32) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
- 33) $\text{Li} + \text{H}_2\text{O} \rightarrow \text{LiOH} + \text{H}_2$
- 34) $\text{Ca}_3(\text{PO}_4)_2 + \text{SiO}_2 + \text{C} \rightarrow \text{CaSiO}_3 + \text{CO} + \text{P}$
- 35) $\text{NH}_3 + \text{O}_2 \rightarrow \text{N}_2 + \text{H}_2\text{O}$
- 36) $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
- 37) $\text{C} + \text{SO}_2 \rightarrow \text{CS}_2 + \text{CO}$

Sequence Chemical Reactions Chemfax Lab Answers

RD Boyd



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The Science Teacher ,2001 **Chemical and Engineering News** ,1963 Rapid Determination of the Order of Chemical Reactions from Time-ratio Tables Joseph H. Flynn,1960 A table of ratios of reaction times corresponding to 15% 5% 20% 10% 95% 85% reaction for orders of chemical reaction from 10 to 10 is presented Use of this table allows the rapid calculation of the order of chemical reactions that are kinetically simple and involves only one subjective step Methods are discussed for interpreting deviations from constant order resulting from errors in the initial times and concentrations from errors in stoichiometry and from the reaction following a more complicated rate expression **Rapid Determination of the Order of Chemical Reactions From Time-Ratio Tables (Classic Reprint)** Joseph H. Flynn,2018-03-29 Excerpt from Rapid Determination of the Order of Chemical Reactions From Time Ratio Tables This paper presents a rapid simple incremental method for determining the order of chemical reactions and may be utilized to detect erroneous assumptions concerning initial conditions and stoichiometry In many cases it provides clues as to the form of complicated kinetic expressions that may be necessary to fit the data Also it minimizes personal bias as the only subjective step involved in the method is the fitting of a smooth curve to a plot of x with respect to time About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books Find more at www.forgottenbooks.com This book is a reproduction of an important historical work Forgotten Books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy In rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition We do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works **Rapid Determination of Order of Chemical Reactions from Time-ratio Tables [with List of Literature Cited]** ,1960

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