

Name: _____

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S'mores Stoichiometry Lab

Introduction:

In this experiment, you will explore the principles of stoichiometry by building S'mores, the delicious, chocolate, marshmallow, and graham cracker treats.



Definitions: 'Stoichio' means element and 'metry' means the process of measuring. The mass and quantity relationships among reactants and products in a reaction are found using the process of stoichiometry.

Purpose:

To determine the limiting reactant in the synthesis of S'mores.

Prelab Questions:

1. What does a chemical equation tell us?
2. What do the coefficients in a balanced chemical equation represent?

Materials:

Miniature Chocolate bars
Marshmallows
Bunsen Burner

Graham cracker Square
Paper Towels
Wooden skewers

Reference Data:

Substance	Symbol	Unit mass	Package Mass
Graham Cracker Squares	Gr	8.50 g	408 g
Marshmallow	Mm	7.125 g	285 g
Miniature Chocolate Bar	Ch	12.00 g	348 g

** for the purposes of this lab the UNIT mass and MOLAR mass are the same thing

Procedure:

1. Perform a synthesis reaction (forming ~~one~~ s'more). For this lab, we will consider a s'more to consist of 2 graham cracker squares, 1 marshmallow and 1 miniature chocolate bar.
 - a) Determine the formula for a s'more
 - b) write the balanced chemical equation using the symbols above
 - c) According to the balanced equation complete the following ratios:

_____ Gr = _____ Mm	_____ Gr = _____ S'more
_____ Ch = _____ Mm	_____ Mm = _____ S'more
_____ Ch = _____ Gr	_____ Ch = _____ S'more
2. Calculating the unit mass of the s'more.
 - a) Calculate the mass on ONE s'more.
 - b) Is there a relationship between the mass of a S'more and the masses of the reactants used to make it?
 - c) If so, what is the relationship?
 - d) What law have you studied in this course that might define this relationship?
3. Predicting the limiting reactant:
 - a) Complete the data table below based on the materials provided for your group.

Substance	Quantity received	# of S'mores it COULD make
Graham Cracker Squares		
Marshmallow		
Miniature Chocolate Bar		
 - b) What will the limiting reactant be?
 - c) How many of the other substances will remain?
 - d) What are the masses of each excess reactant that will remain once the reaction is complete?

Stoichiometry Lab Answer Key

Jennifer Louise Pakkala

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Stoichiometry Lab Answer Key:

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Science Education in the 21st Century Tang Wee Teo, Aik-Ling Tan, Yann Shiou Ong, 2020-06-29 This book reflects on science education in the first 20 years of the 21st century in order to promote academic dialogue on science education from various standpoints and highlights emergent new issues such as education in science education research It also defines new research agendas that should be moved forward and inform new trajectories through the rest of the century Featuring 21 thematically grouped chapters it includes award winning papers and other significant papers that address the theme of the 2018 International Science Education Conference

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