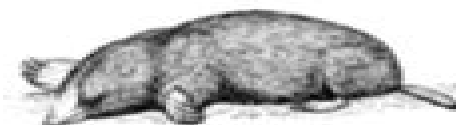


Intro to Stoichiometry – Moles to Moles NOTES

Reminder: To switch between moles of one substance and moles of another, use the ratio of the substances, which is found by looking at their coefficients.

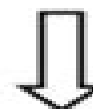


Ex. 1: Given 3 moles of N_2 , how many moles of NH_3 would be produced?



$$\frac{3 \text{ moles } \text{N}_2}{1} \times \frac{2 \text{ moles } \text{NH}_3}{1 \text{ mole } \text{N}_2}$$

Answer: 6.0 moles of NH_3



Ex. 2: Given 0.5 moles of H_2 , how many moles of NH_3 would be produced?



$$\frac{0.5 \text{ moles } \text{H}_2}{1} \times \frac{2 \text{ moles } \text{NH}_3}{3 \text{ mole } \text{H}_2}$$

Answer: 0.33 moles of NH_3

Ex. 3: Given 6 moles of NH_3 , how many moles of N_2 were needed?

$$\frac{6 \text{ moles } \text{NH}_3}{1} \times \frac{1 \text{ moles } \text{N}_2}{2 \text{ mole } \text{NH}_3}$$

Answer: 3.0 moles of N_2

Ex. 4: Given 1.5 moles of N_2 , how many moles of H_2 were needed?

$$\frac{1.5 \text{ moles } \text{N}_2}{1} \times \frac{3 \text{ moles } \text{H}_2}{1 \text{ mole } \text{N}_2}$$

Answer: 4.5 moles of H_2

Stoichiometry Mole Mole Problems Chemistry If8766

John Lambert



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You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 *Bite-Sized Chemistry Calculations* John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume Volume Stoichiometry 13 Percent Yield Answers to Practice Problems Appendices *Solving Problems in Chemistry* Rod O'Connor, Charles Mickey, Alton Hassell, 1977 **Problems and Worked Examples in Chemistry** Albert Holderness, John Lambert, 1978 [A Stoichiometry Unit](#) David Callaghan, 2004

A Tutorial for Solving Mole Problems in Chemistry Bruce Elvert Peffley, 1988 **Stoichiometry and structure** Michell J. Sienko, 1964 [Teaching the Mole Concept in Stoichiometry](#) Timothy Scot Albrecht, California State Polytechnic University, Pomona. College of Education & Integrative Studies, 2009 *Formulation and Stoichiometry* Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the

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Virginia Di-Anna Sutula, 1988

Chemistry Problems Michell J. Sienko, 1967

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