

**KEY**

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of  $K_a$  for the acid HBrO. (4 marks)  $\text{pH} = 4.65$  so  $[\text{H}_3\text{O}^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} \text{ M}$

$$\text{HBrO} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{BrO}^-$$

[I]	0.25		0	0
[C]	$-2.239 \times 10^{-5}$		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		$2.239 \times 10^{-5}$	$2.239 \times 10^{-5}$

(4)

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{BrO}^-]}{[\text{HBrO}]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer  $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks)  $\text{pH} = 2.355$   $[\text{H}_3\text{O}^+] = \text{antilog}(-2.355) = 0.004416 \text{ M}$

$$\text{C}_6\text{H}_5\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{C}_6\text{H}_5\text{COO}^-$$

[I]	$C_0$		0	0
[C]	$-0.004416$		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		$0.004416$	$0.004416$

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_6\text{H}_5\text{COO}^-]}{[\text{C}_6\text{H}_5\text{COOH}]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416$$

$$= 0.30439 \text{ M}$$

Answer  $C_0 = 0.30 \text{ M}$

5. Find the value of  $K_b$  for the oxalate ion ( $\text{C}_2\text{O}_4^{2-}$ ). (1 mark)

$$K_b(\text{C}_2\text{O}_4^{2-}) = \frac{K_w}{K_a(\text{HC}_2\text{O}_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

Answer  $K_b = 1.6 \times 10^{-10}$

6. The value of  $K_b$  for the weak base methylamine ( $\text{CH}_3\text{NH}_2$ ) is  $4.4 \times 10^{-4}$ . Calculate the value of  $K_a$  for the acid  $\text{CH}_3\text{NH}_3^+$ . (1 mark)

$$K_a(\text{CH}_3\text{NH}_3^+) = \frac{K_w}{K_b(\text{CH}_3\text{NH}_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

Answer  $K_a = 2.3 \times 10^{-11}$

# Stoichiometry Notetaking Guide 8answer Key

**M Lipman**



## **Stoichiometry Notetaking Guide 8answer Key:**

## Decoding **Stoichiometry Notetaking Guide 8answer Key**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Stoichiometry Notetaking Guide 8answer Key**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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