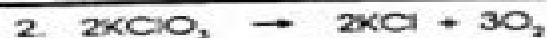


**STOICHIOMETRY:
MOLE-MOLE PROBLEMS**

How many moles of hydrogen are needed to completely react with two moles of nitrogen?

$$\frac{2 \text{ mol N}_2}{1 \text{ mol N}_2} \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} = 6$$

6 moles H₂

How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?

$$6 \text{ moles KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = \frac{18}{2} = 9$$

9 moles of O₂

How many moles of hydrogen are produced from the reaction of three moles of zinc with an excess of hydrochloric acid?

$$3 \text{ mol Zn} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Zn}} = 3$$

3 moles of H₂How many moles of oxygen are necessary to react completely with four moles of propane (C₃H₈)?

$$4 \text{ mol C}_3\text{H}_8 \times \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} = 20$$

20 mol of O₂

How many moles of potassium nitrate are produced when two moles of potassium phosphate react with two moles of aluminum nitrate?

$$\frac{2 \text{ mol K}_3\text{PO}_4}{1 \text{ mol K}_3\text{PO}_4} \times \frac{3 \text{ mol KNO}_3}{1 \text{ mol K}_3\text{PO}_4} = 6$$

6 moles of KNO₃

Stoichiometry Packet Key

Neil D. Jespersen, Pamela Kerrigan



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A Stoichiometry Unit David Callaghan, 2004

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