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# Stoichiometry

Questions and Answers

1. What is the mole ratio in the equation  $2H_2+O_2 \rightarrow 2H_2O$ ?

2:1:2

2. Define stoichiometry.

Stoichiometry is the calculation of reactants and products in chemical reactions based on the balanced equation.

3. How do you determine the limiting reactant?

By calculating the amount of product formed from each reactant and identifying the reactant that forms the least product.

4. What is the theoretical yield?

The maximum amount of product that can be formed from a en reactants.

5. How is percent yield calculated?

Percent Yield=(Actual Yield/R ore at let 100

- 6. Given N<sub>2</sub>+3H<sub>2</sub> + 2MH, w main tes 1 A are produced from 3 moles of H2
- 7. What ar excess pactants?

Reactanks that remain after the limiting reactant is completely used up.

Calculate the theoretical yield if the limiting reactant produces 2 moles of product.

2 moles (converted to grams if needed)

9. Define actual yield.

The amount of product actually obtained from the reaction.

10. If 5 moles of A react with 5 moles of B in the reaction  $A+2B+AB_2$ , which is the limiting reactant?

B

# **Stoichiometry Review Study Guide**

Paul B. Kelter, Steven S. Zumdahl

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